Multicomponent synthesis in water of 7-unsubstituted 4,7-dihydro-
1,2,4-triazolo[1,5-\textit{a}]pyrimidines and their antimicrobial and
antifungal activity

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Abstract
Simple, efficient and eco-friendly method for preparation of 7-unsubstituted 4,7-dihydro-1,2,4-
triazolo[1,5-\textit{a}]pyrimidines was elaborated, which presumes the three-component reaction of 3-
amino-1,2,4-triazole with paraformaldehyde and different 1,3-dicarbonyl compounds in hot-water
by either conventional heating or microwave irradiation without any catalyst. The target
compounds showed antimicrobial and antifungal activity.

Keywords: Multicomponent reaction, heterocycle, 1,4-dihydroazine, hot-water organic synthesis,
microwave-assisted synthesis, green chemistry

Introduction

One of the principles of green chemistry presumes use of safe solvent for preparation of new
substances.\textsuperscript{1} Surprisingly, water, which seems to be the safest possible solvent, was not considered
to be a very suitable reaction medium for organic synthesis during a long time, largely on account
of its non-inert properties and low solubility of most organic compounds in it. Recently the
investigation of organic reactions in water gained some impact. Latest reviews\textsuperscript{2,3} and
monographs\textsuperscript{4,5} showed that use of water as a reaction medium can be useful for carrying out very
different kinds of organic reactions; among them are aldol reaction\textsuperscript{5} and Knoevenagel reaction\textsuperscript{6},
which is especially important because these reactions could be key steps in the multicomponent
synthesis of six-membered heterocycles, such as pyridines and pyrimidines. Although the limited
solubility of most organic compounds in water leads to synthetic limitations, this makes sometimes the isolation of final compound easier, which can be great advantage in some cases.

The main subject of our study are partially hydrogenated azolopyrimidines, compounds which possess interesting biological properties. Among them, 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines are well-studied representatives, the synthetic availability of which is the most developed.

The general method for the preparation of 6,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines includes cyclocondensation reaction of 3-amino-1,2,4-triazole with either \( \alpha,\beta \)-unsaturated carbonyl compound (for example, chalcone derivatives, aryldieneacetones, aryldieneacetoacetates), or with Mannich base of substituted acetophenone as synthetic precursor of \( \alpha,\beta \)-unsaturated carbonyl compound (\textbf{Scheme 1}, Method A).

On the other hand, use of multicomponent reactions is a powerful and efficient tool of modern organic synthesis, which often provides great advantages in comparison with sequential procedures. Therefore, another efficient synthetic approach to 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines is a multicomponent treatment of 3-amino-1,2,4-triazole with aldehyde and methylene-active carbonyl compound. It is presumed, that in such multicomponent reactions the formation of \( \alpha,\beta \)-unsaturated carbonyl compound may occur in the reaction medium. However, some publications have shown that similar treatments proceeded in other ways giving different types of final compounds (\textbf{Scheme 1}, Method B).

\begin{center}
\textbf{Scheme 1.} Two general methods for synthesis of dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines.
\end{center}

There are some examples of multicomponent synthesis of dihydroazolopyrimidine derivatives in water described.

However, in most cases cited above, aryl- or diaryl-substituted derivatives of azoloazines were obtained, and, there are relatively few literature references which describe the preparation of 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines containing no bulky substituent. Unexpectedly, there are only two examples of application of the formaldehyde in such kind of three-component reactions.
The presence of aryl substituent in the molecule of 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidine allows to stabilize the structure of heterocycle, lowers the solubility of compound and makes the isolation process easier. On the other hand, obtaining the azolopyrimidines with lower molecule masses, especially without any aryl ring, can open much wider possibilities for search of new biologically active compounds. It should be noted, that such azolopyrimidines could be used as building-blocks for their chemical modification in order to develop new approaches to wide ranges of compounds based on dihydroazolopyrimidine moiety and with molecular mass of final compound lower than 500 a.u., which is important, taking into account Lipinski principles for biologically active compounds.\textsuperscript{29}

Thus, the main goal of current research was to elaborate efficient and eco-friendly synthetic route for the preparation of novel 7-unsubstituted 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines with no aryl substituent by the three-component reaction of 3-amino-1,2,4-triazoles, paraformaldehyde and 1,3-dicarbonyl compounds, and study of biological activity of target compounds.

**Results and Discussion**

We have established, that the heating an equimolar mixture of 3-amino-1,2,4-triazoles (1\textit{a,b}), paraformaldehyde (2) and 1,3-dicarbonyl compounds 3\textit{a-e} at 100 °C in water without any catalyst leads to the selective formation of 7-unsubstituted 4,7-dihydro-1,2,4-triazolo[1,5-\textit{a}]pyrimidines 4\textit{a-h} in good yields (Scheme 2, Table 1). In case of compound 4\textit{b}, its isolation was possible only after cooling the reaction medium, while in other cases the precipitation of target compounds 4\textit{a,c-f} was observed even during the heating process. In order to increase yields, we attempted to apply microwave irradiation to activate the current reaction. According to our previous experience\textsuperscript{30-35}, microwave-assisted multicomponent reactions are usually more efficient in terms of reaction times, yields and purity in comparison with processes under classical conditions. Surprisingly, our attempts in this case gave the results being similar to those obtained for conventional heating (Table 1).

\[
\text{HNN}^+ - \text{N}^+ - \text{NH}_2 \quad + \quad \text{(HCHO)}_n \quad \overset{\Delta \text{ or MW, H}_2\text{O}}{\rightarrow} \quad R - \text{N}^+ - \text{N}^+ - \text{N}^+ - \text{R}^1 \quad \text{4a-h}
\]

\[
\text{Z} \quad + \quad \text{3a-e} \quad \overset{\Delta \text{ or MW, H}_2\text{O}}{\rightarrow} \quad R - \text{N}^+ - \text{N}^+ - \text{R}^1 \quad \text{or} \quad R - \text{N}^+ - \text{N}^+ - \text{R}^2 \quad \text{4a-h}
\]

\textbf{Scheme 2}. Reaction of 3-amino-1,2,4-triazoles 1 with paraformaldehyde (2) and 1,3-dicarbonyl compounds 3.
Table 1. Used building-blocks 1, 3 and obtained reaction products 4, their yields and reaction conditions

<table>
<thead>
<tr>
<th>Entry</th>
<th>Building-blocks</th>
<th>Reaction product 4</th>
<th>Yield, %</th>
<th>Δ (time, min)</th>
<th>MW (time, min; temperature, °C)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>1a H 3a OC$_2$H$_5$ CH$_3$</td>
<td>4a</td>
<td>72 (15)</td>
<td>75 (15; 100)</td>
<td></td>
</tr>
<tr>
<td>2</td>
<td>1a H 3b CH$_3$ CH$_3$</td>
<td>4b</td>
<td>66 (20)</td>
<td>68 (15; 100)</td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>1a H 3c OCH$_3$ CH$_3$</td>
<td>4c</td>
<td>70 (12)</td>
<td>70 (10; 100)</td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>1b COOCH$_3$ 3a OC$_2$H$_5$ CH$_3$</td>
<td>4d</td>
<td>75 (15)</td>
<td>75 (15; 100)</td>
<td></td>
</tr>
<tr>
<td>5</td>
<td>1a H 3d OC$_2$H$_5$OCH$_3$ CH$_3$</td>
<td>4e</td>
<td>72 (15)</td>
<td>75 (15; 100)</td>
<td></td>
</tr>
<tr>
<td>6</td>
<td>1a H 3e OC$_2$H$_5$ n-C$_3$H$_7$</td>
<td>4f</td>
<td>75 (17)</td>
<td>75 (15; 100)</td>
<td></td>
</tr>
<tr>
<td>7</td>
<td>1b COOCH$_3$ 3b CH$_3$ CH$_3$</td>
<td>4g</td>
<td>72 (15)</td>
<td>75 (13; 100)</td>
<td></td>
</tr>
<tr>
<td>8</td>
<td>1b COOCH$_3$ 3e OC$_2$H$_5$ n-C$_3$H$_7$</td>
<td>4h</td>
<td>68 (20)</td>
<td>70 (20; 100)</td>
<td></td>
</tr>
</tbody>
</table>

The purity and structures of compounds 4a-h were confirmed by their elemental analysis and spectral data ($^1$H and $^{13}$C NMR, IR and MS). Typical $^1$H NMR spectra of all the compounds synthesized showed signals for 6-H methylene protons at 4.79 - 4.95 ppm, broad singlet for NH protons of dihydropyrimidine ring at 10.40 - 10.84 ppm and sharp singlet for 2-H proton of triazole moiety at 7.68-7.71 ppm (for 4a-c,e,f) and other signals for terminal substituents. The $^{13}$C NMR and mass spectra were also in good agreement with the structure proposed. However, it is known,\textsuperscript{22,23,36,37} that the three-component reactions between aminoazoles, carbonyl compounds and CH-acids in some cases can lead to the formation of several isomeric compounds with very similar spectral characteristics. Therefore, in the reaction studied, the isolation of three different heterocyclic compounds may be expected with a certain degree of probability: 4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine (4), 4,5-dihydro-1,2,4-triazolo[1,5-a]pyrimidine (5) or 5,8-dihydro-1,2,4-triazolo[1,5-a]pyrimidine (6). It was shown\textsuperscript{20} that the simplest criterion for assignment of the reaction product either to 4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine or to 4,5-dihydro-1,2,4-triazolo[1,5-a]pyrimidine (structures 4 and 5 respectively) is the chemical shift of NH proton in the $^1$H NMR spectrum, which is about 6.9-7.1 for 4,7-dihydro structure and 9.9-10.0 for 4,5-dihydro structure. Therefore, according to our $^1$H NMR data, all the products 4a-f had structures of 4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine (4), as shown on Scheme 1. Additionally, in the 1D NOE experiment of 4a, the irradiation of NH proton showed enhancement of CH$_3$ proton and no enhancement of CH$_2$ protons; in the NOESY experiment no cross-peaks were observed between CH$_2$ and NH protons and between CH2 and CH proton of triazole moiety, which allowed to reject structures 5 and 6, and therefore, structure of 4 was finally confirmed.
After evaporation of mother liquor no additional amount of 4a was detected; the $^1$H NMR of the residue did not allow to identify signals of heterocycles 4, 5 or 6 and showed only decomposition products.

Obviously, when water is used as a solvent, there is no need in acidic catalyst for the formation of current dihydroazolopyrimidine derivatives, therefore, the mechanism of dihydroazolopyrimidine formation may be different from the classical mechanism of Biginelli reaction$^{38}$ and presumes the previous formation of α,β-unsaturated compound in the Knoevenagel reaction on the first step and followed by heterocyclisation according to Scheme 1 Method A) which includes the 1,4-addition of endocyclic nitrogen to the enone system of unsaturated carbonyl compound. However, it is known$^{39}$ that the preparative formation of such unsaturated carbonyl compound from formaldehyde and acetylacetone is not possible, therefore, the three-component synthesis of azolopyrimidines 4 seems to be the single way here.

Compounds 4 are easily soluble in most organic solvents, like ethanol, and slightly soluble in water, in contrast to their analogs with aromatic rings$^{16}$, therefore, their isolation from reaction media by use other organic solvents can cause some problems.

The compounds 4a,b,d,g,f were tested as substances having antimicrobial and antifungal activity (Escherichia coli ATCC 25922, Staphylococcus aureus ATCC 25923, Pseudomonas aeruginosa ATCC 27853 and Candida albicans ATCC 885-653 strains); as control substances, ampicillin antibiotic and nystatin as antifungal drug were used. The obtained data of minimum inhibitory concentration (MIC), minimum bactericidal or fungicidal concentrations (MBC and MFC respectively) are given in Table 2.

**Table 2. Data of antimicrobial and antifungal activity test for 4a,b,d,f,g**

<table>
<thead>
<tr>
<th>Code name</th>
<th>E.Coli</th>
<th>S.aureus</th>
<th>P.aeruginosa</th>
<th>C.albicans</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>MIC</td>
<td>MBC</td>
<td>MIC</td>
<td>MBC</td>
</tr>
<tr>
<td>4a</td>
<td>100</td>
<td>200</td>
<td>25</td>
<td>50</td>
</tr>
<tr>
<td>4b</td>
<td>100</td>
<td>200</td>
<td>50</td>
<td>100</td>
</tr>
<tr>
<td>4c</td>
<td>100</td>
<td>100</td>
<td>50</td>
<td>100</td>
</tr>
<tr>
<td>4d</td>
<td>100</td>
<td>200</td>
<td>25</td>
<td>50</td>
</tr>
<tr>
<td>4e</td>
<td>100</td>
<td>200</td>
<td>50</td>
<td>100</td>
</tr>
<tr>
<td>4f</td>
<td>100</td>
<td>200</td>
<td>50</td>
<td>100</td>
</tr>
<tr>
<td>4g</td>
<td>100</td>
<td>200</td>
<td>25</td>
<td>50</td>
</tr>
<tr>
<td>Ampicillin</td>
<td>0,78</td>
<td>0,78</td>
<td>0,1</td>
<td>0,195</td>
</tr>
<tr>
<td>Nystatin</td>
<td>100</td>
<td>&gt;200</td>
<td>&gt;200</td>
<td>&gt;200</td>
</tr>
</tbody>
</table>
Our results indicate that all tested compounds exhibit activity against a pathogenic microflora at high MIC, MBC, MFC concentrations in comparison with the control preparation. One of the best results was obtained on compound 4a; ethoxycarbonyl derivative 4a showed a bit higher antimicrobial activity, than acetyl derivative 4b, whereas antifungal activity for 4a and 4b is similar.

**Experimental Section**

**General.** The melting points of all compounds synthesized were determined with a Gallenkamp melting point apparatus. The NMR spectra were recorded at 400 MHz (100 MHz for $^{13}$C) and at 200 MHz (50 MHz for $^{13}$C) with a Varian MR-400 spectrometer. The EI MS spectra were measured on a GC-MS Varian 1200L (ionizing voltage 70 eV, direct input of the sample) instrument. Elemental analysis was realized on EuroVector EA-3000. Analytical samples of the compounds were obtained by their crystallization in water and further drying at room temperature. Microwave experiments were performed using the Emrys Creator EXP from Biotage AB (Uppsala, Sweden) possessing a single-mode microwave cavity producing controlled irradiation at 2.45 GHz. Solvents, all reagents were commercially available and used without additional purification.

The antimicrobial and antifungal activity was carried out on standard Escherichia coli ATCC 25922, Staphylococcus aureus ATCC 25923, Pseudomonas aeruginosa ATCC 27853, Candida albicans ATCC 885-653 strains obtained in bacteriological laboratory of Zaporizhia Regional Laboratory Center of the State Epidemiological Service of Ukraine.

Basic solution was prepared by dissolving 1 mg of investigated substance in 1 mL of dimethylsulfoxide. The 4 mL of Muller-Hinton broth$^{40}$ (for 25923 S. aureus ATCC, E. coli ATCC 25922, P. aeruginosa ATCC 27853) or Sabouraud broth$^{40}$ (for C. albicans ATCC 885-653) were added to basic solution to obtain the dilution with the concentration of substance 200 µg / mL. After series of two-fold dilutions, three additional solutions were prepared with concentrations 100, 50, 25 µg / mL respectively, each in volume of 1 mL. The suspension of microbial culture on amount of 0.1 mL was added into each tube. The test tubes with crops of S. aureus, E. coli, P. aeruginosa were incubated at 37 ± 1 °C for 16-24 hours, with crops of C. albicans - at 28 ± 1 °C for 44-48 hours. MIC was determined on visual absence of bacterial growth in the test tube with minimal concentration of substance investigated.

For determination of MBC and MFC, the content of test tube after MIC determination with no bacterial growth was put into 0.1 mL of Muller-Hinton agar$^{40}$ (for S. aureus, E. coli, P. aeruginosa) or into 0.1 ml of Sabouraud agar$^{40}$ (for C. albicans); cups were incubated at 37 ± 1 °C for 16-24 hours and at 28 ± 1 °C for 44-48 hours respectively. Values of MBC and MFC were detected by visual absence of bacterial growth of cultures.
General procedure for synthesis of 4a-h

Conventional heating. A solution of 3-amino-1,2,4-triazole derivative 1 (0.0012 mol), 0.04 g of paraformaldehyde (2) (0.0013 mol) and 1,3-dicarbonyl compound 3 (0.0012 mol) in water (4.5 mL) was refluxed during the time given in the Table 1. The resulting precipitate was filtered off, recrystallized from water and air-dried.

Microwave activation. An equimolar mixture (0.0012 mol) of 1, 2 and 3 in water (4 mL) was irradiated in MW reactor at 100 °C during the time given it table 1. The crystalline product started to separate out either during the reaction or just after cooling. The precipitate formed was filtered off, washed with water and air-dried.

Ethyl 5-methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-6-carboxylate (4a). White solid, yield 75%, mp 189-190 °C (from water) (Lit.28 194-197°C from ethanol); IR (ν, cm⁻¹): 1711 (C=O), 1665 (C=C). ¹H NMR (400 MHz, DMSO-d₆): δ 1.21 (3H, t, ³J_HH 7.2 Hz, CH₃), 2.29 (3H, s, CH₃), 4.09 (2H, q, CH₂), 4.80 (2H, s, CH₂), 7.70 (1H, s, 2-H), 10.43 (1H, s, NH). ¹³C NMR (100 MHz, DMSO-d₆): δ 13.9, 17.7, 45.2, 59.1, 92.1, 146.6, 147.0, 149.4, 165.0. MS (EI, 70 eV): m/z (%) 208 (44) [M⁺], 180 (19), 179 (100), 163 (10). Anal. Calcd for C₉H₁₂N₄O₂ (208.22): C, 51.92; H, 5.81; N, 26.91%. Found: C, 51.77; H, 5.56; N, 26.63%.

1-(5-Methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidin-6-yl)-ethanone (4b). White solid, yield 68 %, mp 233-234 °C (from water); ¹H NMR (400 MHz, DMSO-d₆): δ 2.21 (3H, s, CH₃), 2.30 (3H, s, CH₃), 4.91 (2H, s, CH₂), 7.72 (1H, s, 2-H), 10.43 (1H, s, NH). ¹³C NMR (100 MHz, DMSO-d₆): δ 19.2, 30.2, 46.0, 102.5, 145.9, 146.9, 149.8, 194.6. MS (EI, 70 eV): m/z (%) 178 (41) [M⁺], 177 (16), 163 (100), 135 (22), 109 (13). Anal. Calcd for C₈H₁₀N₄ (178.19): C, 53.92; H, 5.66; N, 31.44%. Found: C, 54.17; H, 5.90; N, 31.43%.

Methyl 5-methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-6-carboxylate (4c). White solid, yield 70%, mp 241-242 °C (from water); IR (ν, cm⁻¹): 1678 (C=O), 1649 (C=C). ¹H NMR (400 MHz, DMSO-d₆): δ 2.30 (3H, s, 5-CH₃), 3.64 (3H, s, OCH₃), 4.81 (2H, s, CH₂), 7.71 (1H, s, 2-H), 10.47 (1H, s, NH). ¹³C NMR (100 MHz, DMSO-d₆): δ 18.1, 45.4, 51.0, 91.9, 147.1, 147.2, 149.7, 165.7. MS (EI, 70 eV): m/z (%) 194 (41) [M⁺], 179 (100), 161 (25), 163 (10). Anal. Calcd for C₉H₁₀N₄O₂ (194.19): C, 49.48; H, 5.19; N, 28.85%. Found: C, 49.62; H, 5.18; N, 28.67%.

6-Ethyl 2-methyl 5-methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-2,6-dicarboxylate (4d). White solid, yield 75%, 0.35 g, mp 244-245 °C (from water); IR (ν, cm⁻¹): 1740, 1715 (2C=O), 1738 (C=O). ¹H NMR (400 MHz, DMSO-d₆): δ 1.21 (3H, t, ³J_HH 7.2 Hz, CH₃), 2.30 (3H, s, 5-CH₃), 3.81 (3H, s, OCH₃), 4.11 (2H, q, CH₂), 4.86 (2H, s, CH₂), 10.59 (1H, s, NH). ¹³C NMR (100 MHz, DMSO-d₆): δ 18.1, 45.4, 51.0, 91.9, 147.1, 147.2, 149.7, 165.7. MS (EI, 70 eV): m/z (%) 266 (25) [M⁺], 238 (29), 237 (100), 221 (42), 193 (19), 189 (20), 179 (11), 161 (25). Anal. Calcd for C₁₁H₁₄N₄O₄ (266.25): C, 49.62; H, 5.30; N, 21.04%. Found: C, 49.69; H, 5.30; N, 21.11%.

2-Methoxyethyl 5-methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-6-carboxylate (4e). White solid, yield 75%, mp 170-171°C (from water); IR (ν, cm⁻¹): 1700 (C=O), 1645 (C=O). ¹H NMR (400 MHz, DMSO-d₆): δ 2.30 (3H, s, 5-CH₃), 3.27 (3H, s, CH₃), 3.57 (2H, t, CH₂), 4.18
(2H, t, CH2), 4.81 (2H, s, CH2), 7.72 (1H, s, 2-H), 10.50 (1H, s, NH). 13C NMR (100 MHz, DMSO-d6): δ 18.1, 45.3, 58.1, 62.6, 69.9, 91.8, 147.1, 147.4, 149.7, 165.2. MS (EI, 70 eV): m/z (%) 238 (9) [M+], 179 (54), 163 (43), 161 (13), 135 (22), 134 (100). Anal. Calcd for C10H14N4O3 (238.24): C, 50.41; H, 5.92; N, 23.52%. Found: C, 50.49; H, 6.01; N, 23.41%.

**Ethyl 5-propyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-6-carboxylate (4f).** White solid, yield 75%, mp 190-191°C (from water); IR (ν, cm⁻¹): 1705 (C=O). 1H NMR (400 MHz, DMSO-d6): δ 0.91 (3H, t, 3JHH 7.2 Hz, OCH2CH3), 1.21 (3H, s, CH3), 1.56 (2H, m, CH2), 2.69 (2H, t, CH2), 4.10 (2H, q, OCH2CH3), 4.81 (2H, s, CH2), 7.71 (1H, s, 2-H), 10.43 (1H, s, NH). 13C NMR (100 MHz, DMSO-d6): δ 13.7, 14.1, 21.7, 32.6, 45.4, 59.4, 91.7, 147.2, 149.7, 151.1, 165.0. MS (EI, 70 eV): m/z (%) 236 (19) [M+], 208 (19), 207 (100), 149 (11), 134 (19). Anal. Calcd for C11H16N4O2 (236.27): C, 55.92; H, 6.83; N, 23.71%. Found: C, 55.78; H, 6.72; N, 23.84%.

**Methyl 6-acetyl-5-methyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-2-carboxylate (4g).** White solid, yield 75%, mp 248-249°C (from water); 1H NMR (400 MHz, DMSO-d6): δ 2.22 (3H, s, CH3), 2.30 (3H, s, CH3), 3.81 (3H, s, CH3), 4.96 (2H, s, CH2), 10.57 (1H, s, NH). 13C NMR (100 MHz, DMSO-d6): δ 19.2, 30.3, 46.4, 52.2, 103.0, 145.6, 148.0, 151.6, 159.9, 151.6. MS (EI, 70 eV): m/z (%) 236 (48) [M+], 221 (99), 177 (20). Anal. Calcd for C10H12N4O3 (236.23): C, 50.84; H, 5.12; N, 23.72%. Found: C, 50.98; H, 5.26; N, 23.59%.

**6-Ethyl 2-methyl 5-propyl-4,7-dihydro-1,2,4-triazolo[1,5-a]pyrimidine-2,6-dicarboxylate (4h).** White solid, yield 70%, 0.35 g, mp 191-192°C (from water); 1H NMR (400 MHz, DMSO-d6): δ 0.92 (3H, t, CH3), 1.22 (3H, t, 3JHH 7.2 Hz, OCH2CH3), 1.58 (2H, m, CH2), 2.69 (2H, t, CH2), 3.81 (3H, s, CH3), 4.12 (2H, q, OCH2CH3), 4.88 (2H, s, CH2), 10.56 (1H, s, NH). 13C NMR (100 MHz, DMSO-d6): δ 13.6, 14.1, 21.6, 32.6, 45.8, 52.1, 59.6, 92.3, 148.2, 150.7, 151.5, 159.9, 164.8. MS (EI, 70 eV): m/z (%) 294 (9) [M+], 238 (29), 265 (99), 221 (42). Anal. Calcd for C13H18N4O4 (294.31): C, 53.05; H, 6.16; N, 19.04%. Found: C, 53.1; H, 6.11; N, 19.09%.

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