

Novel 1,1-diphenylhydrazine derivatives of benzofurazan and their dimerization

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Abstract

In a quest for novel persistent free radicals, two products were obtained from the reaction between 1,1-diphenylhydrazine and 4-chloro-7-nitrobenzofurazan: the expected 1-(4-nitrobenzofurazan-7-yl)-2,2-diphenyl-hydrazine **3** formed via an S_NAr process, and the unexpected 4,7-benzofurazandione-(bis-1,1-diphenylhydrazone) **4** formed from **3** by a further *ipso* substitution and dehydrogenation. Oxidation of **3** with solid PbO₂ or solid KMnO₄ at room temperature affords during the first few minutes a persistent mono-nitro-hydrazyl free radical **6**, which reacts with NO₂ and forms another persistent hydrazyl free radical **7** having two nitro groups. Two blue dimeric compounds, **11** and **13**, were obtained by dimerisation of **6** and **7** respectively, and loss of N₂O₄, and their structures were determined by mass spectrometry and NMR spectra. Additionally, single crystal X-ray diffraction was used to ascertain the structures

of compounds **4** and **13**. Other novel derivatives were prepared either by *N*-alkylation of **3** with methyl iodide, which afforded compound **5**, or by homolytic nitration of **3** with gaseous NO₂ which yielded the 4,7-dinitrobenzofurazan derivative **8**. When gaseous ¹⁵NNO₂ was used in a similar reaction, only one ¹⁵N-nitro group was present in the product **7**. Persistent free radicals were obtained from **4**, **11**, or **13** with KMnO₄. Cyclic voltammograms for **3** and **8** as well as electronic absorption spectra for all new compounds **3**, **4**, **5**, **8**, **11**, and **13** are discussed.

Keywords: Capto-dative, push-pull, persistent radicals, quinonediimines, cyclic voltammetry, EPR, UV-Vis, RP-TLC, X-ray diffractometry

Introduction

The extreme stability of the free radical 2,2-diphenyl-1-picrylhydrazyl (DPPH)^{1,2} is due to a combination of steric and electronic factors. Steric shielding by two *ortho*-nitro groups prevents its dimerization while the presence of an electron-donor diarylamino group and of a picryl electron-acceptor group provide a unique electronic combination. This “push-pull electronic effect”³ has also been called “capto-dative effect”,⁴ or “merostabilization”.⁵ A quantum-chemical rationalization of such stabilizing effects for free radicals was first provided by Michael Dewar.⁶ Linnett’s double quartet theory⁷ allows a graphical representation of DPPH in terms of spins α and β : the first two resonance structures for the upper row of Chart 1 (with approximately equal contributions as indicated by the quasi-equality of the hyperfine coupling constants of the two nitrogens in the EPR spectrum) correspond to the Linnett structure shown in the lower row, a covalent bond or an unshared electron pair consists of two electrons with opposite spins α and β ; the remaining three electrons with spins α , α and β are distributed between the two nitrogen atoms and their N–N bond. Whereas in the Linnett structure each nitrogen atom is shown to have 8 electrons in the valence shell, the N–N bond has a higher bond order (1.5). With 3,5-di-*tert*-butyl substituents on the phenyl rings and with a benzoyl group replacing the picryl group, it was possible to obtain a simpler EPR spectrum proving by variable-temperature EPR spectroscopy that indeed the rotation barrier around the N–N bond had a higher value in the free radicals than the barrier in corresponding hydrazines.⁸

The present investigation had as initial objective the synthesis of analogs of DPPH having a 4-nitrobenzofurazan electron-withdrawing group, EWG (4-nitro-2,1,3-benzoxadiazole, also abbreviated as NBD) instead of picryl. Such a group would have a lower steric shielding than picryl and a comparable electron-withdrawing effect, and it might impart interesting fluorescence properties. However, interaction of the unpaired spin with an extended chromophoric system could also lead to charge transfer and thus provide a non-radiative decay from the excited state. Like the Boulton-Katritzky rearrangement of nitro-benzofurazan and -benzofuroxan derivatives,⁹

the NBD moiety may also become involved in processes that move the oxadiazole around the benzo ring.¹⁰

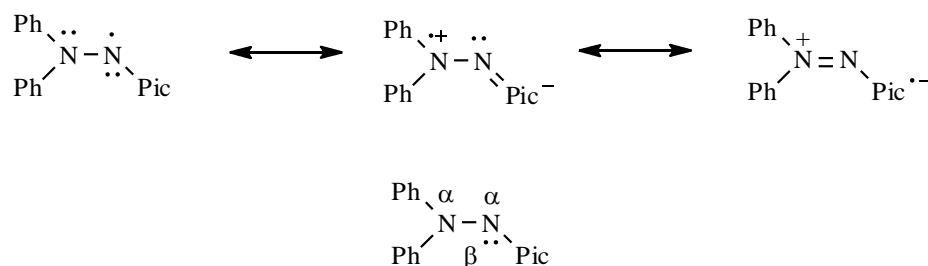


Chart 1. Resonance formulas (upper row) and Linnett structure (lower row) for DPPH (Pic = picryl).

Therefore we performed the reaction between 1,1-diphenylhydrazine **1** and 4-chloro-7-nitrobenzofurazan (**2**, NBD-Cl), obtaining the target push-pull hydrazine **3**, along with an unexpected side-product, **4** (Scheme 1). We then investigated in detail chemical and physical properties of the novel compounds which, in spite of having one or two NBD moieties, proved to be non-fluorescent.

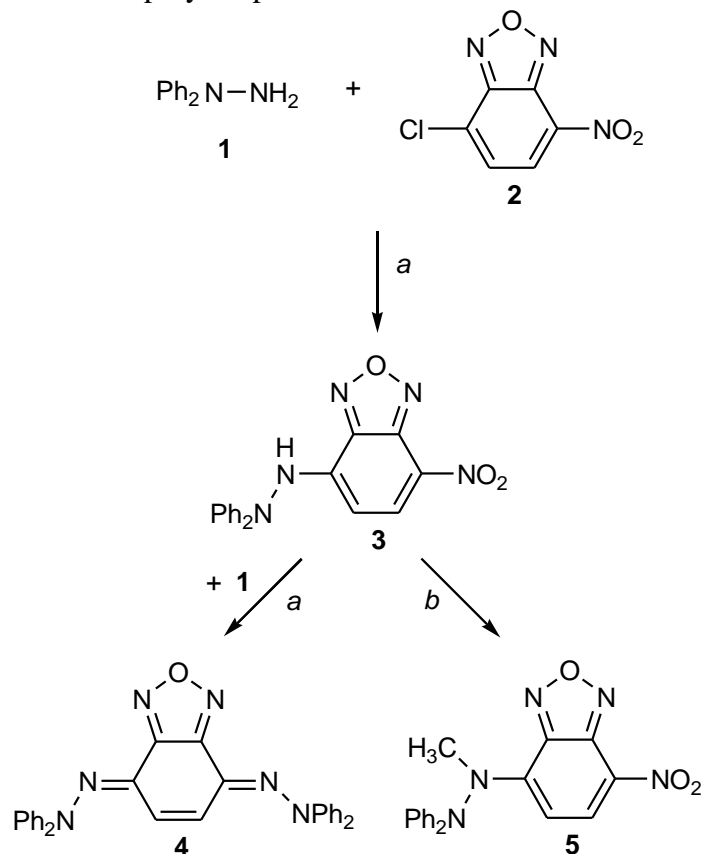
Reaction of 4-chloro-7-nitrobenzofurazan with 1,1-diphenylhydrazine and structures of the products

1,1-Diphenylhydrazine **1** was stirred in ethanol at room temperature with an excess of sodium hydrogen carbonate to which an equimolar amount of NBD-Cl was added. After heating at reflux for one hour and work-up of the reaction mixture, two products were obtained by preparative thin layer chromatography (TLC): 1-(4-nitrobenzofurazan-7-yl)-2,2-diphenylhydrazine **3** in 30% yield, and 4,7-benzofurazandione-(bis-1,1-diphenylhydrazone) **4** in 21% yield. The reaction mechanism for the expected product **3** involves an S_NAr process,¹¹ while for the rather unexpected product **4** an *ipso*-substitution of the nitro group by the nucleophilic **1** must be invoked followed by an oxidation to form the quinonediimine system (Scheme 1). Indeed, on treating separately 1,1-diphenylhydrazine **1** with the push-pull NBD-diphenylhydrazine **3** in warm ethanol in the presence of $NaHCO_3$, we obtained again compound **4** but only in 18% isolated yield.

By reacting compound **3** with methyl iodide under conditions of phase transfer catalysis,¹² the *N*-methyl-derivative **5** was obtained in 52% yield after TLC, confirming the structure of **3**. In contrast to the highly reactive **3** or **4**, the air-stable *N*-methyl derivative **5** did not show multiple spots upon TLC purification.

Single crystals of **4** could be grown by slow diffusion of *n*-heptane into dry dichloromethane, and interestingly two polymorphs were encountered: (i) rather large orange-ochre prisms which could be easily cleaved and (ii) very thin but rather long red needles. The red needles were mounted at the ANKA Synchrotron in Karlsruhe, and although they diffracted X-rays well, giving a reasonable data set, their structure could not be solved by direct methods. In

contrast, the large prisms could be solved by conventional techniques showing without doubt that **4** was present in this mixture of polymorphs.



Scheme 1. Formation of compounds **3** and **4** from the reaction of 1,1-diphenylhydrazine (**1**) and NBD-Cl (**2**). Reagents and conditions: *a*) NaHCO₃, EtOH, Δ , 1 h; *b*) CH₃I, NaOH, 15-crown-5, CH₂Cl₂, RT, 5 h.

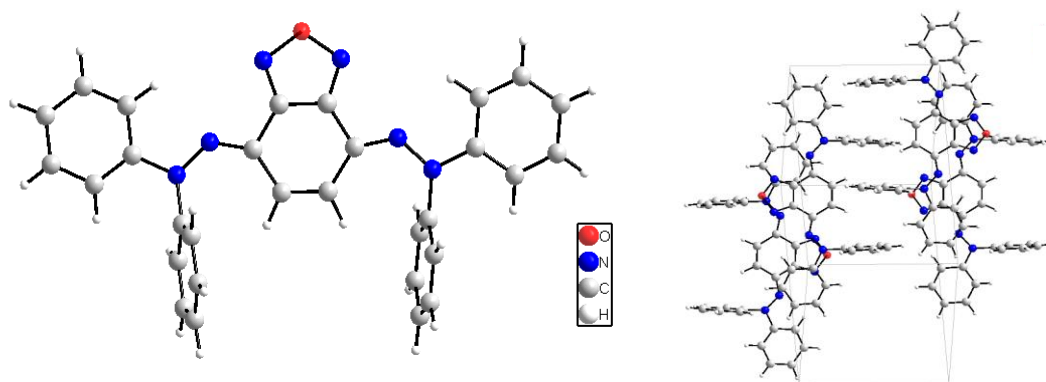


Figure 1. Structure of an orange single crystal of **4** as determined by X-ray analysis. At right, the crystal packing showing four molecules within the unit cell.

In the crystal, the furazan O–N bonds are 1.389 Å long with an NON angle of 111.6°. The exocyclic quinonoid C=N double bonds are 1.301 Å long and tilted slightly towards the five-membered ring aligning almost collinearly the bonds of the two *para*- nitrogen atoms on either side of the C_{2v} symmetric molecule. Interestingly, concerning these quinonoid double bonds, only the (*E,E*) isomer shown in Figure 1 has been found, which can be explained by a probable electrostatic repulsion between the furazan nitrogen atom and the sp^3 -hybridized hydrazonic nitrogen atom in hypothetical *Z*-stereoisomers (but possibly the red needles may contain another stereoisomer). The two *N*-phenyl bonds are slightly different in length, one being 1.417 Å while the other one is 1.440 Å long, a fact reflecting that the former phenyl ring is engaged in strong π – π interactions (Figure 1 and S1 – Supplementary Material). The furazan 5-membered ring overlaps completely with such a phenyl ring ensuring a 2-D connected network with 3.367 Å as the shortest intermolecular distance. Additionally, as shown in the right part of Figure 1, the other *N*-phenyl groups stack in a perpendicular direction forming a staircase-like architecture with 3.45 Å shortest distances between the ladders.

Formation of free radicals from compound 3

The electron-withdrawing 4-nitrobenzofurazan substituent and the diphenylamino group make compound **3** a typical precursor for persistent push-pull aminyl radicals related to DPPH. On treating compound **3** in deaerated methylene chloride with solid PbO_2 or solid $KMnO_4$ at room temperature, two free radicals were evidenced by EPR spectroscopy, as shown in Figure 2.

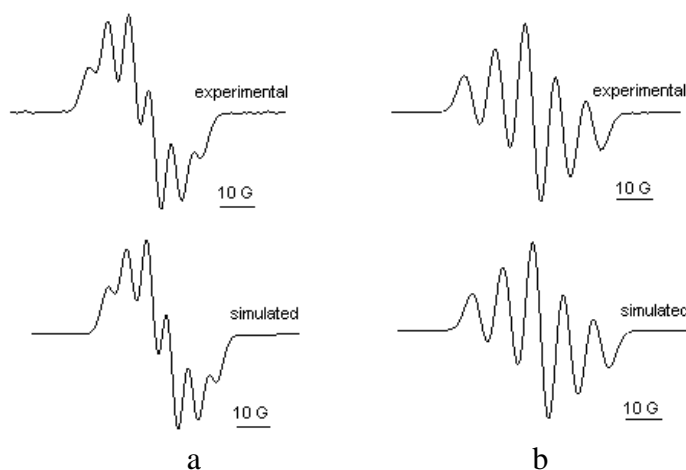
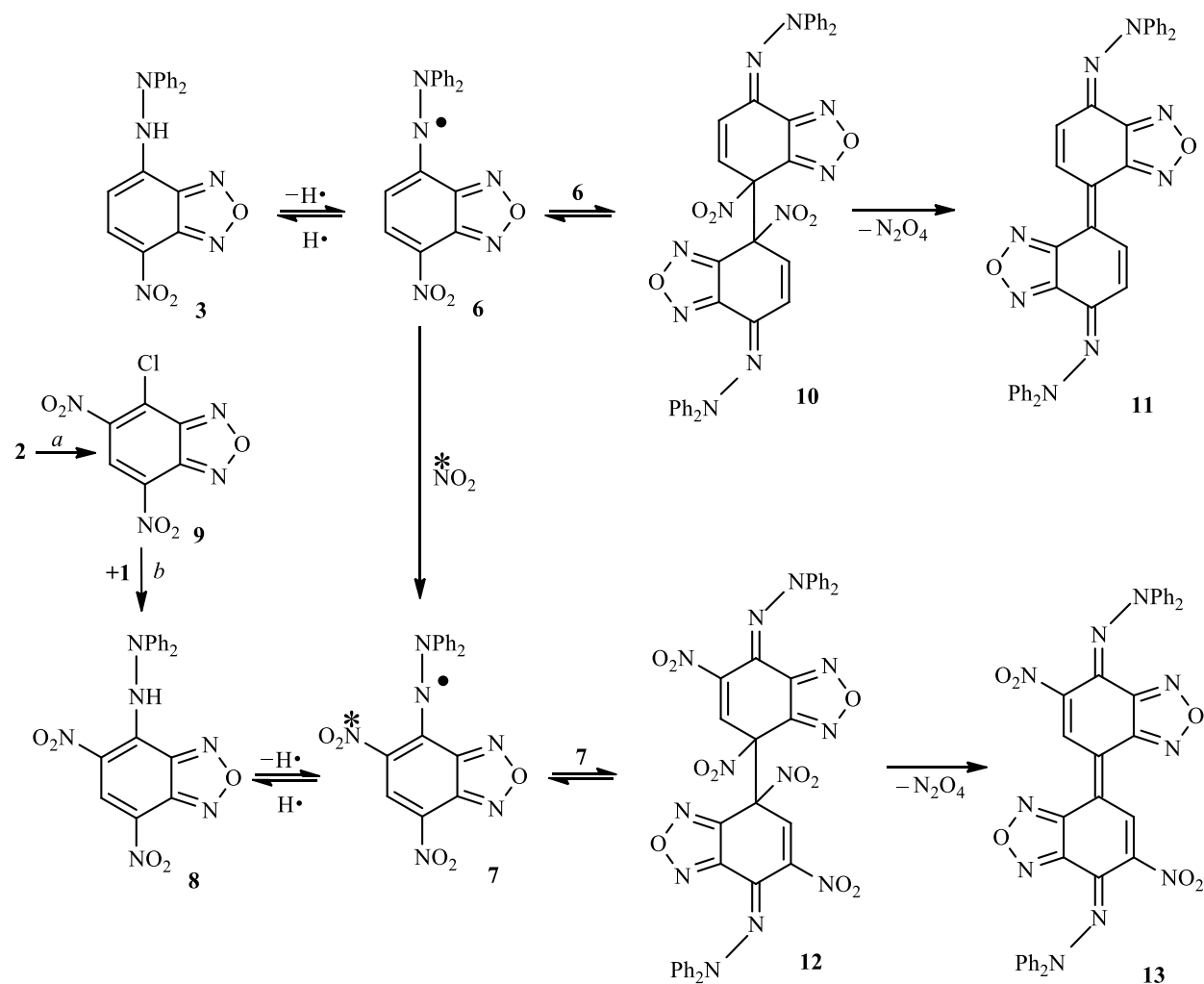


Figure 2. Experimental (upper traces) and simulated (lower traces) of EPR spectra of oxidation products of **3** in CH_2Cl_2 with solid $KMnO_4$ after 3 minutes (a) and after 7 minutes (b).

By preparative TLC analysis of the reaction mixture after 7 – 10 minutes, the new spots, along with unreacted **3**, allowed the detection and identification of compounds **11** and **13**, formed

from the corresponding persistent free radicals **6** and **7**, respectively, according to the reaction mechanism depicted in Scheme 2.



Scheme 2. Formation of compounds **8**, **11**, and **13** via the persistent free radicals **6** and **7** as well as the not isolated intermediates **10** and **12**. $^{*}\text{N}$ in $^{*}\text{NO}_2$ stands for either ^{14}N or ^{15}N . Reagents and conditions: *a*) HNO₃/oleum, 100-105 °C; *b*) NaHCO₃, EtOH, Δ , 1 h.

The oxidation of **3** leads to the initial formation of radical **6** displaying an EPR spectrum with 6 lines, which does not have the steric shielding of DPPH, and therefore dimerizes easily to **10** which then disproportionates affording $2\text{NO}_2 \rightleftharpoons \text{N}_2\text{O}_4$ and the deep-blue diphenoquinonediiminic compound **11** exhibiting in the mass spectrum a molecular peak at $m/z = 600$ Daltons. Concurrently, **6** also reacts with the *in situ* formed NO_2 as shown in Scheme 2 affording at slightly longer reaction times the free radical **7** having an EPR spectrum with 5 lines. In analogous reactions to preceding ones, dimerization and disproportionation of **7** leads to a different blue compound, **13** with the molecular peak at $m/z = 690$ Daltons. Reversible hydrogen

abstraction from **3** transforms radical **7** into the neutral dinitro-substituted compound **8**. One should stress that structural assignments for **6** and **7** are based only on plausible interpretations of ESR spectra.

In a separate experiment we could confirm the homolytic nitration of **3** with gaseous NO₂.¹³ Nitrogen dioxide, which is more reactive than nitric oxide, was generated from solid powdered sodium nitrite and a solid support,¹⁴ namely silica gel for chromatography, in a biphasic solid – liquid system mediated by the crown ether 15-crown-5 in methylene chloride. We also used separately ¹⁵N-labeled sodium nitrite, and the asterisk in Scheme 2 can signify either one of the two stable nitrogen isotopes. It was thus found that the dinitro-derivative **8** is formed by homolytic substitution in 35-40% yield. Thus, 1-(4,6-dinitrobenzofurazan-7-yl)-2,2-diphenylhydrazine was formed when using normal NaNO₂, and 1-(4-[¹⁴N]nitro-6-[¹⁵N]nitro-benzofurazan-7-yl)-2,2-diphenylhydrazine was obtained when using [¹⁵N]NaNO₂. No scrambling of the label occurred.

An alternative synthesis of **8** could be achieved by first nitrating NBD-Cl (**2**) with nitric acid and oleum¹⁵ followed by conversion of the resulting dinitro-derivative **9** into **8** by an *S_NAr* process on using diphenylhydrazine (**1**) as the nucleophile. In this case no *ipso*-substitution of either nitro group could be observed.

For both free radicals **6** and **7**, the *g*-factor agrees with the value associated with hydrazyls.² Hyperfine coupling constants used for the simulated EPR spectra are presented in Table 1.

Table 1. EPR data for hydrazyls **6** and **7** (Scheme 2; Figure 2a for **6** and Figure 2b for **7**)

Free radical	<i>g</i>	Hyperfine coupling constants
		(Gauss)
6	2.0037	<i>a</i> _{N1} = 6.04, <i>a</i> _{N2} = 5.54, <i>a</i> _{N3} = 3.14, <i>a</i> _{N4} = 1.00
7	2.0037	<i>a</i> _{N1} = 8.04, <i>a</i> _{N2} = 8.03

A further experiment confirming the structural assignments was the oxidation of **3** at room temperature in CH₂Cl₂ with solid KMnO₄ in the presence of ascorbic acid (molar ratio 10:1 for **3**:ascorbic acid). The EPR signal of **6** as in Figure 2a was constant for about 15 minutes after which it decreased, but no formation of **7** as in Figure 2b was observed. On the other hand, oxidation of compound **8** with solid KMnO₄ at room temperature in CH₂Cl₂ afforded radical **7** with five lines as in Figure 2b (intensity ratio 1:2:3:2:1), *g* = 2.0037, *a*_{N1} = 8.15 and *a*_{N2} = 8.05. The slightly different *a_N* values from those presented in Table 1 are probably due to the fact that the latter values are affected by small amounts of **6** present in solution.

In the case of DPPH, no dimerization is detected in solution (although in single crystals some evidence for dimer formation was reported).¹⁶ Definite proof for the dimeric structure of compound **13** came again from single crystal X-ray diffraction. Using similar conditions as for the growth of crystals of **4**, we obtained thin blue needles that required synchrotron radiation for their X-ray analysis. The structure was solved by direct methods showing that among the six possible *syn-anti* stereoisomers, the predominant form is the one presented in Figure 3.

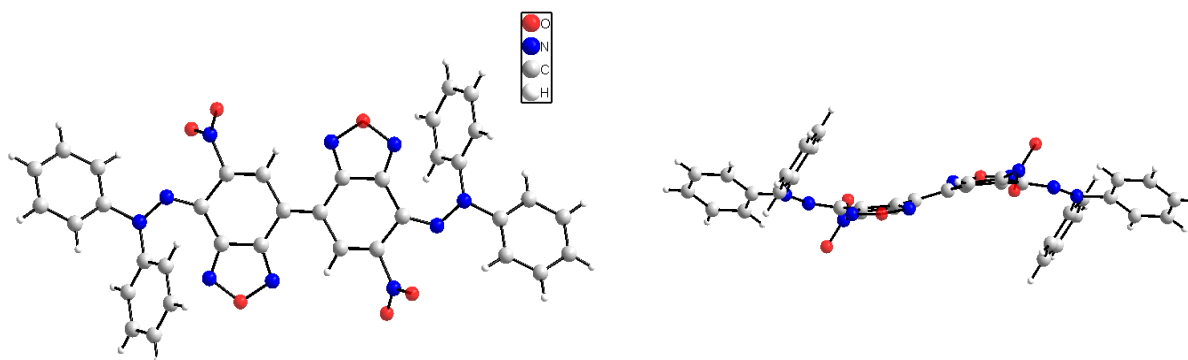


Figure 3. Structure of **13** ($m/z = 690$) by single crystal X-ray analysis. At right, a view of the nonplanar conjugated system.

In structure **13** the two benzofurazan moieties are oriented in opposite directions with a common central C=C double bond which is 1.385 Å long. This particular double bond is thus *trans* and enables a weak C-H...N hydrogen bond between the phenyl protons and the furazan nitrogen atoms. A *cis* double bond would have positioned two furazanic nitrogens in close proximity generating a less stable structure. The two central benzo rings in **13** are coplanar to within less than 2°. In the case of a central single bond the biphenyl system would not have been coplanar exhibiting a rotation barrier giving rise to enantiomers. Interestingly, chirality is however present in the crystal due to the S-shaped form adopted by the conjugated system, as shown in the right part of Figure 3. Extended chromophoric systems involving diphenoquinonediimides were thoroughly investigated more than 50 years ago by Roger Adams and coworkers.¹⁷ In our particular case, the combination of the furazan ring and the nitro group on the benzo ring, contrary to **4**, ensures that in **13** both diphenylamino moieties of the exocyclic diphenoquinonediimide are now facing the furazan ring. Formally they are still *E* as the nitro group has a higher priority than the furazan nitrogen atom. Thus the correct designation of the obtained stereoisomer **13** is *E-trans-E*. The nitro groups have a 52.9° torsion angle to the phenyl rings and all point in the same direction within stacks of off-setted molecules having a 5.48 Å distance between neighboring nitrogen atoms of the nitro group. The C–N distance is 1.478 Å, typical for non-conjugated nitroarenes. The C=N bonds are rather short (1.309 Å) while the N–N bond is 1.330 Å. Packing diagrams of compound **13** in the crystal are presented in Figure S2 – Supplementary Material.

Formation of persistent radicals from compounds (**4**, **11** and **13**)

Interestingly, compounds **4**, **11** and **13** present as solids weak but reproducible broad EPR signals. Furthermore, when submitted to TLC, although initially pure, several new spots appear upon elution of the silica gel plates. No rigorous exclusion of oxygen was made, neither from the solvents nor from the elution chamber, so that these bis-hydrazones might actually be readily

oxidized to corresponding mono-hydrazyl-*N*-oxides. Due to the absence of a captodative stabilization and absence of severe steric hindrance, their persistence is much lower than that of **6**, **7** or DPPH.

Compound **4** in the solid state presents an EPR spectrum with $g = 2.0031$. On dissolution in CH_2Cl_2 this signal disappears and reappears on evaporating the solvent (Figure 4a). By oxidizing compound **4** with solid KMnO_4 in CH_2Cl_2 a persistent free radical (presumably the radical-cation from **4**) with five lines is obtained (Figure 4b) with the following hyperfine coupling constants: $a_{\text{N1}} = 8.50 \text{ G}$, $a_{\text{N2}} = 8.48 \text{ G}$ and with $g = 2.0034$. The intensity of the EPR spectrum decreases appreciably in 10 minutes.

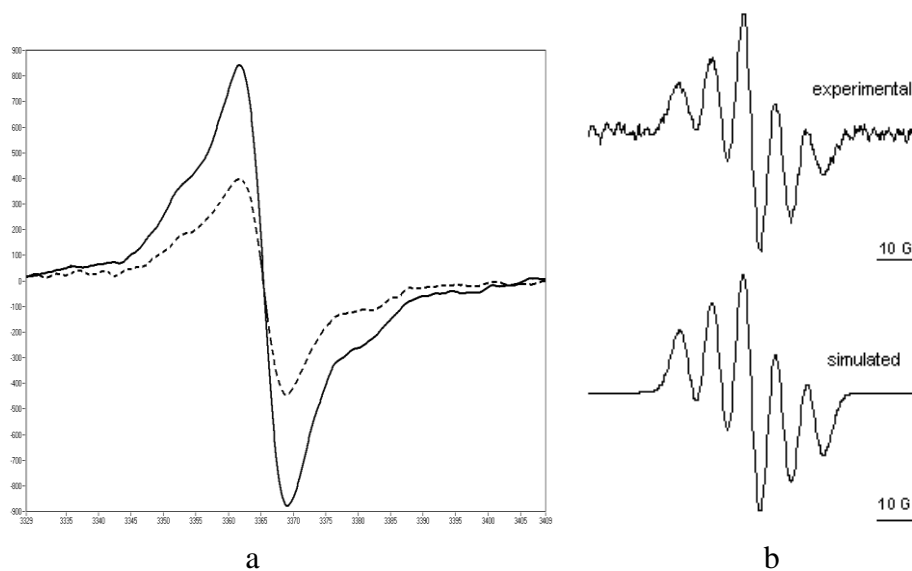


Figure 4. EPR Spectrum of compound **4**; a, in solid state (full line); after dissolution in methylene chloride and evaporation of the solvent (dashed line); b, after dissolution in methylene chloride and oxidation with KMnO_4 solid (experimental and simulated).

A similar behaviour is encountered for the dimeric compounds **11** and **13** which both exhibit in the solid state a weak EPR signal. Upon dissolution in methylene chloride this signal disappears but it is reformed upon evaporation of the solvent. By oxidizing compounds **11** and **13** with solid KMnO_4 in CH_2Cl_2 , persistent free radicals (presumably radical-cations) with five-line ESR spectra are obtained (Figures 5a, 5b, and Table 2); their intensity decreases in time, faster for the less sterically-shielded structure devoid of nitro groups.

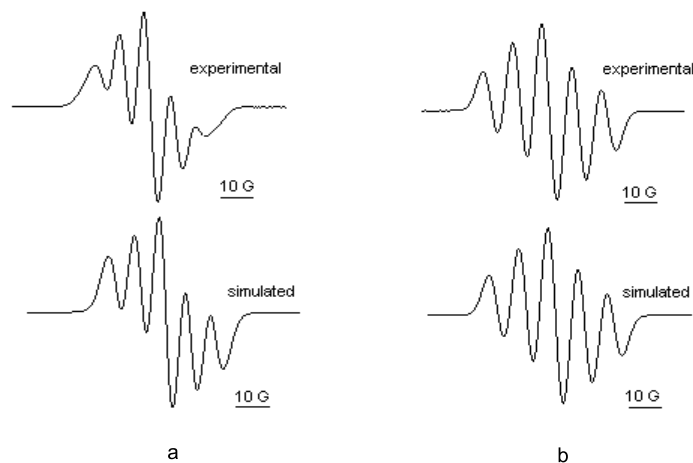


Figure 5. EPR Spectra in deaerated methylene chloride after oxidation with solid KMnO_4 : a, for **11** (experimental and simulated) and b, for **13** (experimental and simulated)

Table 2. EPR Spectra of the radicals obtained by oxidizing dimeric compounds **11** and **13** with solid KMnO_4 in CH_2Cl_2 (Figure 5)

Dimeric compound	g	Hyperfine coupling constants (Gauss)
11	2.0033	$a_{\text{N1}} = 7.10$, $a_{\text{N2}} = 7.07$
13	2.0033	$a_{\text{N1}} = 9.21$, $a_{\text{N2}} = 8.28$

Cyclic voltammetry

Comparative studies were undertaken using cyclic voltammetry for three related compounds (2,2-diphenyl-1-EWG-hydrazines) with electron-withdrawing groups (EWGs): **3**, **8**, and DPPH-H with 1, 2, and 3 nitro groups, respectively. The results are presented in Figures 6 and 7, and Table 3. While compound **3** is oxidized at the same potential as DPPH-H but it is reduced only partly, compound **8** is oxidized and reduced at higher potentials than **3** and DPPH-H.

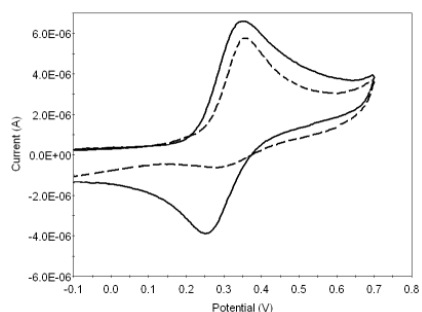


Figure 6. Cyclic voltammograms recorded for a concentration of 10^{-3} M of DPPH-H (full line) and **3** (dashed line) in acetonitrile (vs. Ag/AgCl) at a scan rate of 0.5 V s^{-1} ; supporting electrolyte 10^{-1} M TBAP (tetra-*n*-butylammonium perchlorate); scanning potential from -0.1 V to 0.7 V.

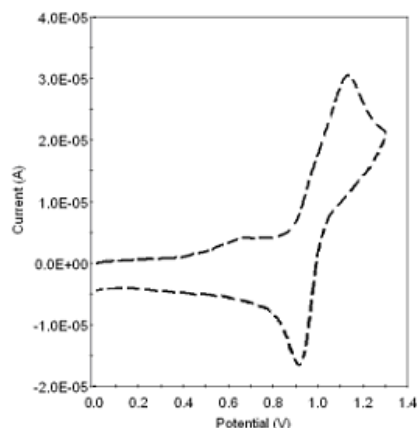


Figure 7. Cyclic voltammogram for **8** at a concentration of 10^{-3} M; scan rate of 0.1 V s^{-1} , supporting electrolyte 10^{-1} M TBAP; scanning potential from 0 V to 1.3 V (vs. Ag/AgCl).

Table 3. Peak potentials (V vs. Ag/Ag⁺) for oxidation (E_{pa}I) and reduction (E_{pc}I) of DPPH-H, **3** and **8** in acetonitrile in the presence of TBAP (10^{-1} M) (Figures 6 and 7).

Compound	E _{pa} I	E _{pc} I
DPPH-H	0.353	0.250
3	0.353	0.292
8	1.13	0.920

The cyclic voltammogram of compound **4** is shown in Figure 8 and it has waves corresponding to the oxidized form **4**, the reduced form with two NH bonds, and also a peak with low potential value that is probably due to a partly reduced free radical with one NH bond (Table 4).

Absorption spectra of compounds (**3**, **4**, **5**, **8**, **11**, **13**) and the anions of (**3**) and DPPH-H

Table 5 lists the UV-Vis absorption bands of DPPH-H and **3**, of their anions, and of compounds **4**, **5**, **8**, **11**, and **13**. One can observe that compounds **3**, **8**, and the anion of **3** absorb at appreciably higher wavelengths than DPPH and DPPH[−] (anion of DPPH-H). This bathochromic effect must be due to the difference between the NBD moiety and the picryl group. It should be mentioned again that none of the new compounds (**3**, **4**, **5**, **8**, **11**, or **13**) were fluorescent.

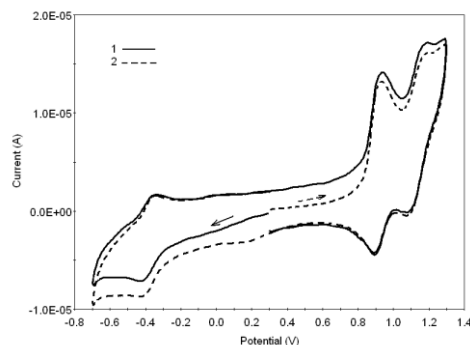


Figure 8. Cyclic voltammogram for a concentration of 10^{-3} M of compound **4** at a scan rate of 0.1 V s^{-1} at 30°C ; supporting electrolyte: 10^{-1} M TBAP. Full line (1), scanning potential from 0.3 V to -0.7 V: first (cathodic) step, from -0.7 V to 1.3 V: second (anodic) step; and from 1.3 to 0.3 V: third (backward) step. Dashed line (2), scanning potential from 0.3 V to 1.3 V: first (anodic) step; from 1.3 V to -0.7 V: second (cathodic) step; and from -0.7 V to 0.3 V: third (backward) step (all potentials vs. Ag/AgCl).

Table 4. Oxidation and reduction peak potentials (V vs. Ag/AgCl) of compound **4** in the presence of TBAP as supporting electrolyte (concentration, 10^{-1} M)

1 st peak		2 nd peak		3 rd peak		ΔEpI^a	ΔEpII^b	ΔEpIII^c
EpaI	EpcI	EpaII	EpcII	EpaIII	EpcIII			
-0.360	-0.420	0.940	0.885	1.170	1.095	0.060	0.055	0.080

^a $\Delta\text{EpI} = \text{EpaI} - \text{EpcI}$; ^b $\Delta\text{EpII} = \text{EpaII} - \text{EpcII}$; ^c $\Delta\text{EpIII} = \text{EpaIII} - \text{EpcIII}$

Table 5. Absorption maxima (λ_{max}) in the visible region in CH_2Cl_2 for DPPH-H, its anion (DPPH^-), compounds **3**, **4**, **5**, **8**, **11**, **13** and the anion of **3**

Compound	λ_{max} (nm)	ϵ ($\text{L} \times \text{M}^{-1} \times \text{cm}^{-1}$)
DPPH-H	318	13100
DPPH^- ^a	424	-
3	451	9500
Anion of 3 ^a	405	16800
	500	28210
4	483	Not determined
5	332	6230
	477	13900
8	389	17800
	492	8370
11	636	Not determined
13	486	Not determined
	617	Not determined

^aDPPH-H and **3** in the presence of triethylamine.

Hydrophobic-hydrophilic balance

The hydrophobicity of organic compounds is of prime importance for their applications, especially in biomedicine. The experimental partition coefficient between 1-octanol and water ($\log P$) can be calculated fairly accurately by means of various programs, e. g. $\text{Alog}P$.¹⁸⁻²⁰ Another experimental approach is based on reverse-phase TLC (RP-TLC) and produces the molecular hydrophobicity, R_{M0} .²¹⁻²³ In the following we present this parameter for the newly synthesized compounds **3**, **4**, and **5** and for the oxidation products of **3**, namely **8**, **11**, and **13**. For comparison we also include DPPH-H, and the starting materials **1** (1,1-diphenylhydrazine), **2** (NBD-Cl), **9** (dinitro-benzofurazan-Cl), and picryl chloride (**14**). Experiments with RP-TLC involved the stationary phase C_{18} covalently bonded to silica gel, and various concentrations of aqueous ethanol as the mobile phase. The molecular hydrophobicity, R_{M0} was calculated by means of equations (1) and (2), where the alcohol concentration K was extrapolated to zero, and b is the slope of the plot for the R_M value when the alcohol concentration is varied from A = 95% to D = 80%, as seen in Table 6. The same Table presents the statistical parameters R^2 (the correlation coefficient R is always a negative number close to -1), Fisher parameter F , and the standard deviation SD .

$$R_M = \log(1/R_f - 1) \quad (1)$$

$$R_M = R_{M0} + bK \quad (2)$$

Table 6. Comparative results for R_M , R_{M0} , b and calculated $\text{Alog} P$ of compounds **1**, **2**, **3**, **4**, **5**, **8**, **9**, **11**, **13**, **14** and DPPH-H ^a

Comp.	R_M				Hydrophobicity characteristics		Statistical parameters			$\text{Alog}P$ calc.
	A	B	C	D	R_{M0}	b	R^2	F	SD	
1	-0.698	-0.423	-0.309	-0.128	2.803	-0.036	0.972	69.19	0.049	5.97
2	-0.562	-0.465	-0.408	-0.34	0.822	-0.014	0.987	153.61	0.013	1.96
3	-0.616	-0.434	-0.309	-0.154	2.266	-0.030	0.995	407.19	0.016	4.48
4 ^(b)	-0.239	0	0.176	0.421	3.863	-0.043	0.996	528.10	0.021	7.70
5	-0.483	-0.265	-0.128	0.043	2.793	-0.034	0.992	242.37	0.024	4.84
8	-0.635	-0.503	-0.382	-0.196	2.087	-0.029	0.990	203.17	0.022	4.38
9	-0.602	-0.527	-0.479	-0.423	0.516	-0.012	0.991	225.89	0.009	1.85
11 ^(b)	0.239	0.524	0.736	1.010	5.046	-0.051	0.997	677.18	0.021	7.91
13 ^(b)	0.383	0.673	0.857	1.085	4.757	-0.046	0.992	251.39	0.032	8.43
14	-0.717	-0.562	-0.454	-0.36	1.541	-0.024	0.986	141.16	0.022	2.18
DPPH-H	-0.434	-0.252	-0.106	0.042	2.567	-0.031	0.997	685.90	0.013	4.70

^aFour determinations on silica gel RP-18F_{254S} (Merck), with ethanol concentrations in ethanol-water mixture: A = 95%, B = 90%, C = 85% and D = 80%; ^(b) For major spot (from oxidized or reduced forms).

One can see from Table 6 that the most hydrophobic compounds are **11** and **13**, and the least hydrophobic are the starting chlorinated materials. The molecular hydrophobicity decreases in the following order: R_{M0} **11** > R_{M0} **13** > R_{M0} **4** > R_{M0} **1** > R_{M0} **5** > R_{M0} DPPH-H > R_{M0} **3** > R_{M0} **8** > R_{M0} **14** > R_{M0} **2** > R_{M0} **9**. A plot of the experimental R_{M0} values *versus* the calculated $AlogP$ partition coefficient^{19,20} (Figure 9) presents a satisfactory linear correlation, as indicated in Table 7 and in equation 3.

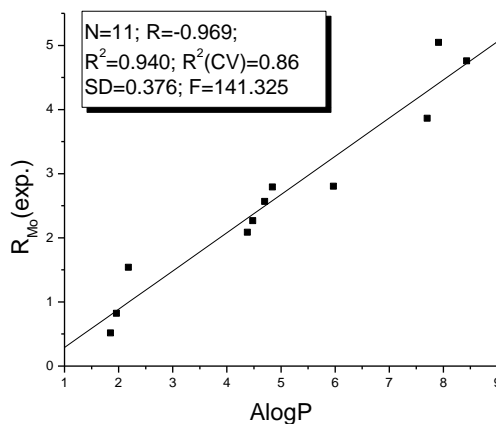


Figure 9. R_{M0} vs $AlogP$ for compounds **1**, **2**, **3**, **4**, **5**, **8**, **9**, **11**, **13**, **14** and DPPH-H.

$$R_{M0}(\text{calcd.}) = 0.596(\pm 0.050)AlogP - 0.306 \quad (3)$$

Table 7. Comparison between R_{M0} (exp.) and R_{M0} (calcd. from $AlogP$ by equation 3)

Compound	$AlogP$	R_{M0}		
		Exp.	Calcd.	Residual
1	5.97	2.803	3.252	0.449
2	1.96	0.822	0.862	0.040
3	4.48	2.266	2.364	0.098
4 ^a	7.70	3.863	4.283	0.420
5	4.84	2.793	2.578	0.215
8	4.38	2.087	2.304	0.217
9	1.85	0.516	0.796	0.280
11 ^a	7.91	5.046	4.408	0.638
13 ^a	8.43	4.757	4.718	0.039
14	2.18	1.541	0.993	0.548
DPPH-H	4.70	2.567	2.495	0.072

^a For major spot (from oxidized or reduced forms)

Conclusions

The reaction between 1,1-diphenylhydrazine (**1**) and NBD-Cl (**2**) in the presence of NaHCO₃ proceeds normally by an S_NAr mechanism affording 1-(4-nitrobenzofurazan-7-yl)-2,2-diphenylhydrazine (**3**), but then this compound reacts further with **1** by *ipso*-substitution of its nitro group yielding 4,7-benzofurazandione-(bis-1,1-diphenylhydrazone) (**4**), whose structure was proven by IR and NMR spectra, *N*-methylation to **5**, and by X-ray diffraction of single crystals. A dinitro congener of **3**, namely 1-(5,7-dinitrobenzofurazan-4-yl)-2,2-diphenylhydrazine (**8**) resulted both by homolytic nitration of **3** with gaseous NO₂, and by heterolytic S_NAr attack of **1** on chloro-dinitro-benzoxadiazole (DNBD-Cl, **9**). Like the trinitro compound DPPH-H, the dinitro derivative **8** does not react with **1**, whereas the mono-nitro congener **3** does. These three 2,2-diphenyl-1-EWG-hydrazines, namely 2,2-diphenyl-1-picrylhydrazine (DPPH-H), **3**, and **8**, form persistent hydrazyls on oxidation, but unlike DPPH which does not dimerize in solution, **3** and **8** do. The unstable dimers **3** and **8** evolve 2NO₂ \rightleftharpoons N₂O₄ and yield stable blue compounds **11** and **13**, respectively. The structure of **13** was firmly established by X-ray diffraction of single crystals using synchrotron radiation.

Redox processes of **3**, **4** and **8** were investigated chemically and by cyclic voltammetry. Like DPPH-H, the two congeners **3** and **8** react reversibly with bases such as tertiary amines forming stable anions. Compounds **3** and **8** absorb visible light at longer wavelengths than DPPH-H, and the quinone-diiminic structures **4**, **11** and **13** exhibit even stronger bathochromic effects. On oxidation with KMnO₄, compounds **4**, **11** and **13** yield persistent free radicals, presumably radical-cations.

Acknowledgements

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Experimental Section

General. Commercial chemicals were purchased: from Merck, *N,N*-diphenylhydrazinium chloride (**1**-HCl), 15-crown-5, methyl iodide, TLC glass plates with silica gel 60 F₂₅₄ (for analytical TLC), silica gel 60 PLC-F₂₅₄ (for preparative TLC), silica gel RP-18 F₂₅₄S (for RP-TLC) and silica gel 60 (0.063 – 0.200 mm) for column chromatography (1.07734 code); from Fluka, DPPH-H, tetra-*n*-butylammonium perchlorate and ammonium sulfamate; and from Aldrich, 4-chloro-7-nitrobenzofurazan (**2**); from the Institute of Stable Isotopes Cluj Romania,

Na¹⁵NO₂ (98.3% enriched). Compound **9**¹⁵ and picryl chloride **14**²⁴ were prepared according to literature data.

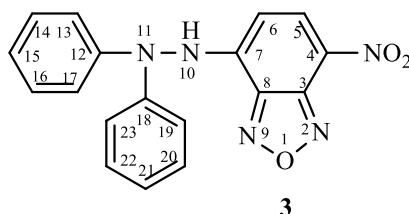
Electronic absorption spectra were recorded with a Unicam-UV-VIS spectrophotometer using *Vision Software V.3.33*. NMR spectra were recorded with a Varian Gemini 300 BB spectrometer (300 MHz for ¹H-NMR and 75 MHz for ¹³C-NMR) using TMS as internal standard. Fourier-transform IR spectra were recorded on a Bruker Vertex 70 spectrometer with horizontal device for attenuated reflectance and diamond crystal, on a spectral window ranging from 4000 to 400 cm⁻¹. EPR spectra were recorded with a Jeol JES-FA 100 instrument. Mass spectra were recorded with Varian 1200 L/MS/MS Triple Quadrupole with ESI interface both positive and negative ionization, fragmentation by collision with argon at 1.5 mTorr. Melting points have been recorded in open capillary with Electrothermal IA 9000 Series of digital melting point instruments. Cyclic voltammetric measurements were performed with a conventional three-electrode glass cell by means of a PAR-273-A potentiostat, and all solutions were prepared by using acetonitrile. As reference electrode, a Ag/AgCl electrode immersed in a 0.1 M supporting electrolyte solution was used, linked to the main compartment of the cell by means of a Vycor plug. A platinum disk (surface area, 0.07 cm²) and a platinum wire were used as the working and counter electrode, respectively. As supporting electrolyte, tetra-*n*-butylammonium perchlorate (TBAP) was used at a concentration of 0.1 M.

It should be stressed that the atom numberings indicated below are only for assistance with NMR assignments.

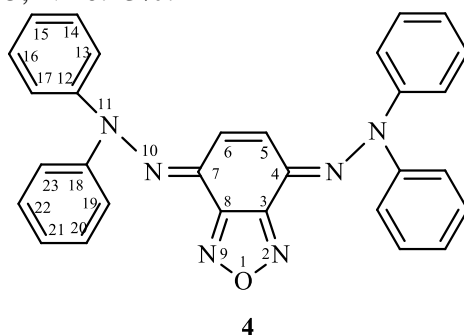
1-(4-Nitrobenzofurazan-7-yl)-2,2-diphenyl-hydrazine (3) and 4,7-benzofurazan-dione-(bis-1,1-diphenylhydrazone) or 6,6'-dinitrobibenzofurazan-4,4'-ylidene-7,7'-dione bis-(2,2-diphenylhydrazone) (4). General procedure starting from (1) and (2)

1,1-Diphenylhydrazine **1** and sodium hydrogen carbonate (molar ratio 1 : 2.5) were stirred for 15 min. at room temperature in ethanol (50 mL / 1g of **1**), then NBD-Cl (**2**) was added (equimolar ratio **1** : **2**) and the mixture was refluxed for one hour. The brown-violet solution was diluted with water till a precipitate was formed; the mixture was acidified with 1N hydrochloric acid to pH = 1; liquid-liquid extraction with methylene chloride followed till the organic extract was colorless. The separated and dried organic solvent was evaporated under reduced pressure, and the residue was dried in a desiccator over CaCl₂. The products were separated and purified by preparative TLC (silica gel 60 PLC-F₂₅₄, with dichloromethane : toluene = 1 : 1 v/v, R_f **3** = 0.25, yield 30% and R_f **4** = 0.575, yield 21%).

Atom numberings in the following formulas are *ad hoc* assistance for NMR assignments and do not follow IUPAC rules.



1-(4-Nitrobenzofurazan-7-yl)-2,2-diphenylhydrazine (3). Brick-red crystals, m.p. 148-149°C. ESI-MS, (m/z); for $C_{18}H_{13}N_5O_3$ (**3**, $M=347$), in negative and positive: $[M-H]^-$ 346 and $[M+H]^+$ 348. FT-IR (solid ATR, ν cm^{-1}): 994m, 1272s, 1307vs, 1367m, 1441s, 1492s, 1556w, 1586m, 1752w, 2852m, 2922s, 3295s. 1H -NMR ($CDCl_3$, δ ppm, J Hz): 8.46(d, 1H, H-5, 8.6), 8.06(bs, 1H, NH, deuterable), 7.35(dd, 4H, H-14, H-16 and H-20, H-22, 8.4, 7.4), 7.20(dd, 4H, H-13, H-17 and H-19, H-23, 1.3, 8.4), 7.15(tt, 2H, H-15 and H-21, 1.3, 7.4), 6.78(d, 1H, H-6, 8.6). ^{13}C -NMR ($CDCl_3$, δ ppm): 144.97(C-12 and C-18), 143.93(Cq-1), 142.70(C-3), 142.25(Cq), 135.77(C-5), 130.07(Cq), 129.82(C-14, C-16 and C-20, C-22), 124.71(C-15 and C-21), 119.70(C-13, C-17 and C-19, C-23), 110.87(C-6). $C_{18}H_{13}N_5O_3$ (347.31): calcd. C 62.25, H 3.77, N 20.16; found C 62.23, H 3.75, N 20.13%.

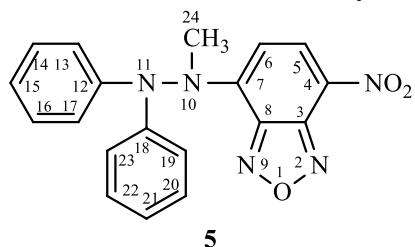


4,7-Benzofurazandione-(bis-1,1-diphenylhydrazone) (4). Deep-red crystals with metallic luster, mixture containing small amounts of reduced form (no precise m.p. or elemental analysis). ESI-MS, m/z ; for mixture $C_{30}H_{22}N_6O$ (**4**, $M=482$) and $C_{30}H_{24}N_6O$ (**14**, $M=484$), in positive: $[M]^+$ 482, $[M+H]^+$ 483, $[M]^+$ 484, $[M+H]^+$ 485. FT-IR (ATR in solid, ν cm^{-1}): 3059w, 2921w, 2852w, 1586m, 1521s, 1480vs, 1396m, 1330m, 1203vs, 1130m, 1078w, 1037s, 993m, 946m, 871w, 788w, 740m, 691m, 607w, 582w, 542w. 1H -NMR($CDCl_3$, δ ppm, J Hz): 7.36(dd, 8H, H-14, H-16 and H-20, H-22, 8.2, 7.4), 7.26(tt, 4H, H-15 and H-21, 1.4, 7.4), 7.20(dd, 8H, H-13, H-17 and H-19, H-23, 1.4, 8.2), 5.42(s, 2H, H-5, H-6). ^{13}C -NMR ($CDCl_3$, δ ppm): 150.11(C-4, C-7), 146.28(C-12, C-18), 129.77(C-14, C-16 and C-20, C-22), 128.91(Cq), 125.88(C-15 and C-21), 122.51(C-13, C-17 and C-19, C-23), 119.24(C-5, C-6). Note that the atom numbering for the NMR assignments differs from the one in the deposited CIF file CCDC-812719.

4,7-Benzofurazandione-(bis-1,1-diphenylhydrazone) (4). General procedure starting from compounds (1) and (3)

1,1-Diphenylhydrazine (**1**) and sodium hydrogen carbonate (molar ratio 1 : 2) were stirred for 15 min. at room temperature in ethanol (50 mL/1g of **1**), then compound **3** was added (equimolar

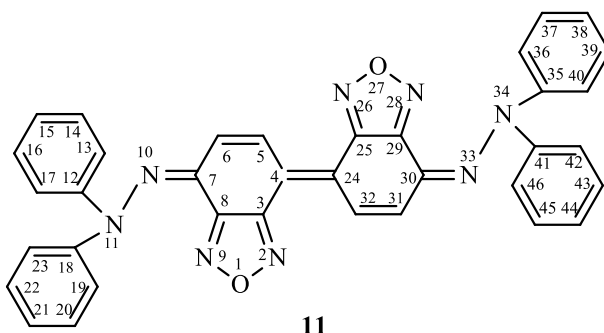
ratio) and the mixture was refluxed for one hour. The red-violet reaction mixture was worked up as described above. The red product **4** was obtained in 18% yield.



1-Methyl-1-(4-nitrobenzofurazan-7-yl)-2,2-diphenylhydrazine (5). Under phase-transfer catalytic conditions by stirring in methylene chloride with compound **3**, solid powdered NaOH, 15-crown-5 and CH₃I in molar ratios **3** : NaOH : 15-Crown-5 : CH₃I = 1 : 1 : 1.5 : 6 as follows: For 20 mL/0.5 g reactant mixture, the freshly ground NaOH was suspended in CH₂Cl₂ then 15-crown-5 was added and the mixture was stirred till complete dissolution occurred. Then **3** and CH₃I were added and the mixture was stirred at room temperature for 5 hrs, when the red color turned yellow-orange. After removal of the solvent under vacuum, the product **5** was isolated and purified by preparative TLC (Silica Gel 60 PLC-F₂₅₄, methylene chloride : toluene = 1 : 1 v/v, R_f **5** = 0.375). Yield 52% in yellow-orange crystals, m.p. 164.5-165°C: ESI-MS, *m/z*; for C₁₉H₁₅N₅O₃ (**5**, M=361), in positive: [M+H]⁺ 362. FT-IR (ATR in solid, ν cm⁻¹): 3234w, 3044w, 2926w, 2855w; 1620m; 1587m; 1546vs; 1491vs; 1460m; 1448m; 1422m; 1369w; 1311vs; 1291vs, 1270vs, 1204m, 1175m, 1133w, 1107w, 1082m, 1042w, 1031w, 1003m, 977w, 914m, 866w, 825w, 779w, 750s, 739m, 693m, 630w, 616w. ¹H-NMR (CDCl₃, δ ppm, *J* Hz): 8.41(d, 1H, H-5, 9.0), 7.34(dd, 4H, H-14, H-16 and H-20, H-22, 8.4, 7.4), 7.14(dd, 4H, H-13, H-17 and H-19, H-23, 1.3, 8.4); 7.11(tt, 2H, H-15 and H-21, 1.3, 7.4); 6.75 (bs, 1H, H-6), 4.01(bs, 3H, H-24), broadened signals for H-6 and C-6 due to hindered rotation caused by the methyl group. ¹³C-NMR(CDCl₃, δ ppm): 145.44(Cq), 144.77(C-12 and C-18), 143.87(C-1), 142.26(C-3), 135.05(C-5), 129.88(C-14, C-16 and C-20, C-22), 129.66(Cq), 124.34(C-15 and C-21), 119.21(C-13, C-17 and C-19, C-23), 102.93(bs, C-6); 40.72(C-24). C₁₉H₁₅N₅O₃ (361.34): calcd. C 63.15, H 4.18, N 19.38; found C 63.13, H 4.16, N 19.35%.

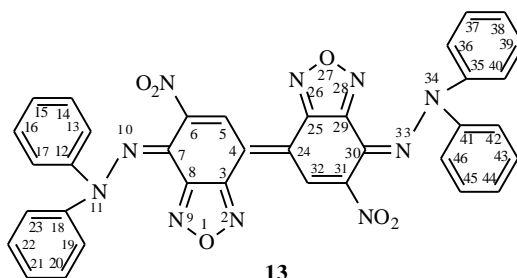
7,7'-Bis-benzofurazanylidene-4,4'-bis-(2,2-diphenylhydrazone) (11), 1-(4,6-dinitrobenzofurazan-7-yl)-2,2-diphenylhydrazine (8), and 7,7'-bis-5,5'-dinitrobenzofurazanylidene-4,4'-bis-(2,2-diphenylhydrazone) (13). General procedure by oxidation of (**3**)

An excess of finely ground potassium permanganate was added to a solution of **3** in CH₂Cl₂ (0.5 g KMnO₄ / 1 g of **3**). After stirring for 24 hrs at room temperature the mixture was centrifuged, retaining the supernatant solution and the solid was extracted three times with CH₂Cl₂. The combined solutions were concentrated in vacuum, and the three products were separated by preparative TLC (Silica Gel 60 PLC-F₂₅₄, eluent methylene chloride : toluene = 1:1 v/v): a red-violet zone for **8** (R_f = 0.475, yield 5%), a blue zone for **11** (R_f = 0.706, yield 14%), and a blue zone for **13** (R_f = 0.800, yield 11%).



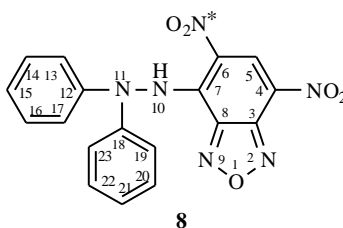
11

7,7'-Bis-benzofurazanylidene-4,4'-bis-(2,2-diphenylhydrazone) (11) Deep blue solid with metallic luster, mixture containing small amounts of reduced form (no precise m.p. or elemental analysis). ESI-MS, m/z ; for mixture $C_{36}H_{24}N_8O_2$ (**11**, $M=600$) and $C_{36}H_{26}N_8O_2$ (**18**, $M=602$), in positive: $[M]^{+}$ 600, $[M+H]^{+}$ 601, $[M]^{+}$ 602. FT-IR(ATR in solid, ν cm^{-1}): 2924vs, 2854s, 1734w, 1590w, 1480s, 1449s, 1413w, 1370m, 1313m, 1276m, 1219m, 1184m, 1085w, 1056w, 999w, 971w, 888w, 749m, 693m, 580w. 1H -NMR ($CDCl_3$, δ ppm, J Hz): 8.87(s, 1H, NH), 8.66(s, 1H, NH), 8.20(d, 2H, H-5, H-32, 9.3), 7.30 (d, 2H, H-6, H-31, 9.3), 7.48÷7.15 (m, 20H, H-arom). Doublet peaks at 8.10 (H-5) and 5.86 ppm (H-6) with $J = 10.5$ Hz confirms presence of a reduced form in a tenfold lower amount. ^{13}C -NMR ($CDCl_3$, δ ppm): 151.58(Cq), 148.02(Cq), 144.49(Cq), 141.43(Cq), 129.39(C-14, C-16 and C-20, C-22), 124.96(C-13, C-17 and C-19, C-23 or C-36, C-40 and C-42, C-46), 123.78(C-13, C-17 and C-19, C-23 or C-36, C-40 and C-42, C-46), 122.66(C-15, C-21, C-38, C-44), 119.80(C-5, C-32 or C-6, C-31), 119.62(C-5, C-32 or C-6, C-31).



13

7,7'-Bis-5,5'-dinitro-benzofurazanylidene-4,4'-bis-(2,2-diphenylhydrazone) (13). Deep blue solid with metallic luster, mixture containing small amounts of reduced form (no precise m.p. or elemental analysis). ESI-MS, m/z ; for mixture $C_{36}H_{22}N_{10}O_6$ (**13**, $M=690$) and $C_{36}H_{24}N_{10}O_6$ (**19**, $M=692$), in positive: $[M]^{+}$ 690, $[M+H]^{+}$ 691, $[M]^{+}$ 692, $[M+H]^{+}$ 693. FT-IR (solid ATR), cm^{-1} : 963m, 1204vs, 1305m, 1366m, 1451m, 1480vs, 1589s, 1725w, 2856m, 2923s. 1H -NMR ($CDCl_3$, δ ppm, J Hz): 8.70(s, 2H, H-5), 7.37(dd, 4H, H-14, H-16 and H-20, H-22, 8.4, 7.2), 7.31(tt, 2H, H-15 and H-21, 1.5, 7.2), 7.23(dd, 4H, H-13, H-17 and H-19, H-23, 1.5, 8.4). ^{13}C -NMR ($CDCl_3$, δ ppm): 144.68(C-12, C-18), 129.31(C-14, C-16 and C-20, C-22), 127.68(C-15 and C-21), 123.66(C-13, C-17 and C-19, C-23), 120.57(C-5), 120.35(Cq), 152.26(C-4, C-7), 147.96(Cq), 142.03(Cq-3), 129.35(Cq). Note that the atom numbering for the NMR assignments differs from the one in the deposited CIF file CCDC-812720.



1-(4,6-Dinitro-benzofurazan-7-yl)-2,2-diphenylhydrazine (8). Violet crystals, m.p. 184-186°C. ESI-MS, m/z ; for $C_{18}H_{12}N_6O_5$ (**8**, $M=392$), in negative: $[M-H]^-$ 391. FT-IR(ATR in solid, ν cm^{-1}): 3226m, 3066w, 2907w, 1625s, 1573s, 1525s, 1485s, 1415m, 1378m, 1299vs, 1269vs, 1180s, 1125s, 1072m, 1027m, 1000m, 938m, 892w, 854w, 814w, 776w, 738m, 687m, 650m, 577w. 1H -NMR ($CDCl_3$, δ ppm, J Hz): 11.73(s, 1H, H-N, deuterable), 9.36(s, 1H, H-5), 7.36(dd, 4H, H-*meta*, H-14, H-17 and H-20, H-22, 7.5, 8.4), 7.21(dd, 4H, H-*ortho*, H-13, H-17 and H-19, H-23, 1.3, 8.4), 7.18(tt, 2H, H-*para*, H-15 and H-21, 1.3, 7.5). ^{13}C -NMR ($CDCl_3$, δ ppm): 144.68(C-12 and C-18), 143.97(C-1), 143.68(C-3), 137.13(C-4), 129.86(C-14, C-16 and C-20, C-22), 129.77(C-8), 129.69(C-5), 125.52(C-15 and C-21), 123.29(d, C-6), 120.35(C-13, C-17 and C-19, C-23). $C_{18}H_{12}N_6O_5$ (392.31): calcd. C 55.10, H 3.08, N 21.42; found C 55.08, H 3.05, N 21.39%.

1-(4,6-Dinitro-benzofurazan-7-yl)-2,2-diphenylhydrazine and its $^{15}NO_2$ -labeled congener (8). **General procedure starting from 3 by homolytic nitration.** A solution of 30 mg $NaNO_2$ or $[^{15}N]NaNO_2$ in 25 mL CH_2Cl_2 was obtained with the help of 15-crown-5 (molar ratio $NaNO_2$: crown ether = 1 : 1.5). Compound **3** was added to this solution (molar ratio **3** : $NaNO_2$ = 1 : 1.5) together with solid silica gel for column chromatography (ratio 0.5 g silicagel / 0.1 g of **3**). The mixture was stirred for 24 hrs, monitoring the course of the reaction by TLC analysis (Silica Gel 60 F₂₅₄, methylene chloride : toluene = 1 : 1 v/v) by following the decrease of **3** (R_f = 0.25) and the formation of compound **8** (R_f = 0.475). The mixture was centrifuged, retaining the supernatant solution, and the solid was extracted three times with CH_2Cl_2 . The combined solutions were extracted twice with water with ammonium sulfamate, once with dilute hydrochloric acid with ammonium sulfamate, and finally again with water. The organic layer was dried over Na_2SO_4 , the solvent was removed in vacuum, and the violet product was purified by preparative TLC (Silica Gel 60 PLC-F₂₅₄, CH_2Cl_2 : toluene = 1 : 1 v/v), yield 40%.

1-(4-[^{14}N]Nitro-6-[^{15}N]nitro-benzofurazan-7-yl)-2,2-diphenylhydrazine, **8- ^{15}N .** ESI-MS, m/z ; for $C_{18}H_{12}^*N_6O_5$ (**8** - ^{15}N , $M = 393$), in negative: $[M-H]^-$ 392. FT-IR (ATR in solid, ν cm^{-1}): 3225m, 3065w, 2923m, 2855w, 1625s, 1570s, 1523s, 1487s, 1460s, 1415m, 1373m, 1295vs, 1264vs, 1175s, 1124s, 1073m, 1000m, 936m, 891w, 814w, 776w, 743m, 686m, 649m, 577w, 547w. 1H -NMR ($CDCl_3$, δ ppm, J Hz): 11.73(s, 1H, H-N, deuterable), 9.36(d, 1H, H-5, $^3J(H^5-^{15}N)=1.9$ Hz), 7.36(dd, 4H, H-*meta*, H-14, H-17 and H-20, H-22, 7.3, 8.5), 7.21(dd, 4H, H-*ortho*, H-13, H-17 and H-19, H-23, 1.3, 8.5), 7.18(tt, 2H, H-*para*, H-15 and H-21, 1.3, 7.3). ^{13}C -NMR ($CDCl_3$, δ ppm): 144.68(C-12 and C-18), 143.96(d, C-1, $^2J(C^1-^{15}N)=2.7$ Hz), 143.67(C-3), 137.13(C-4), 129.86(C-14, C-16 and C-20, C-22), 129.77(C-8), 129.69(d, C-5, $^2J(C^5-^{15}N)=2.1$

Hz), 125.53(C-15 and C-21), 123.29(d, C-6, $J(\text{C}^6\text{-}^{15}\text{N}) = 17.9$ Hz), 120.34(C-13, C-17 and C-19, C-23). Elemental analysis $\text{C}_{18}\text{H}_{12}\text{N}_6\text{O}_5$ (393.31): calcd. C 54.96, H 3.07, N 21.62; found C 54.94, H 3.05, N 21.59%.

1-(5,7-Dinitro-benzofurazan-4-yl)-2,2-diphenylhydrazine (8). General procedure starting from NBD-Cl (2)

Nitration of 4-chloro-7-nitrobenzofurazan **2** (NBD-Cl or 4-chloro-7-nitro-2,1,3-benzoxadiazole) was carried out according to literature data.¹⁵ The dinitro-derivative **9** was separated from unreacted **2** by preparative TLC in 6% yield (Silica Gel 60 PLC-F₂₅₄, CH_2Cl_2 : methanol = 9 : 1 v/v, R_f **2** = 1.0, R_f **9** = 0.34); m.p. for **9** found 143.5-144.5 °C (lit.¹⁵ 143-144 °C). For the synthesis of **8** from **9**, the procedure followed the operations indicated for obtaining **3** from **2**; TLC monitoring showed that the reaction was completed in 15 minutes. The yield was 60% based on **9**.

X-Ray analysis

Crystals suitable for single crystal X-ray diffraction were selected in perfluoroalkyl-ether oil. Single-crystal X-ray diffraction data of **4** were collected with graphite-monochromatized $\text{MoK}\alpha$ radiation ($\lambda = 0.71073$ Å) on a STOE IPDS II (Imaging Plate Diffraction System). Single-crystal X-ray diffraction data of **13** were collected using synchrotron radiation ($\lambda = 0.80$ Å) on a STOE IPDS II at the ANKA synchrotron source in Karlsruhe. Raw intensity data were collected and treated with the STOE X-Area software Version 1.27. Data were corrected for Lorentz and polarisation effects while the dataset of **13** was also corrected for the intensity decay of the beam. The structures were solved with the direct methods program SHELXS of the SHELXTL PC suite programs,²⁵ and were refined with the use of the full-matrix least-squares program SHELXL.²⁵ Molecular diagrams were prepared using Diamond.²⁶ All C, O, and N atoms were refined with anisotropic displacement parameters. All H atoms could be localized in the difference Fourier map and were refined without restraints. Table 8 summarizes the main data collection and refinement parameters.

Cambridge Database documents CCDC-812719 (**4**) and CCDC-812720 (**13**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge at www.ccdc.cam.ac.uk/conts/retrieving.html (or from the Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: (internat.) +44-1223/336-033; E-mail: deposit@ccdc.cam.ac.uk). In the Supplementary Material one can find packing diagrams for **4** and **13**.

Table 8. Crystallographic data for C₃₀H₂₂N₆O (**4**) and C₃₆H₂₂N₁₀O₆ (**13**)

		4	13
<i>fw</i> [g mol ⁻¹]		482.54	690.64
crystal system		monoclinic	triclinic
space group		<i>C</i> 2 <i>c</i>	<i>P</i> $\bar{1}$
Cell	<i>a</i> [Å]	20.014(4)	5.483(1)
	<i>b</i> [Å]	9.5335(19)	9.621(2)
	<i>c</i> [Å]	12.576(3)	15.613(3)
	α [°]		106.46(3)
	β [°]	99.53(3)	94.81(3)
	γ [°]		91.88(3)
	<i>V</i> [Å ³]	2366.4(8)	785.7(3)
<i>Z</i>		4	1
<i>T</i> [K]		180	150
λ [Å]		Mo–K α	0.80000
<i>d_c</i> [g cm ⁻³]		1.354	1.460
$\mu(\lambda)$ [mm ⁻¹]		0.086	0.129
<i>F</i> [000]		1008	356
2 θ_{\max} [°]		54	54
measured reflexions		8774	7413
unique reflexions		2475	2301
<i>R</i> _{int}		0.0750	0.0937
reflexions with <i>I</i> > 2 $\sigma(I)$		2234	1930
refined parameters		212	279
<i>R</i> 1 (<i>I</i> > 2 $\sigma(I)$) ^a		0.0425	0.0526
<i>wR</i> 2 (all data) ^b		0.1361	0.1485

$$^a R1 = \sum ||F_o| - |F_c|| / \sum |F_o|. \quad ^b wR2 = \{ \sum [w(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2] \}^{1/2}$$

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