

Reaction of ketene and [dimethyl(phenyl)silyl]ketene with the aminoxyl radical TEMPO: preparative, kinetic, and theoretical studies of unreactive ketenes

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Dedicated to Joseph Muchowski on the occasion of his 65th birthday, and in recognition of his achievements in heterocyclic chemistry in Canada, Mexico, and the United States

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Abstract

Ketenes $\text{CH}_2=\text{C}=\text{O}$ (**1**) and $\text{PhMe}_2\text{SiCH}=\text{C}=\text{O}$ (**12**) react with tetramethylpiperidinyloxy (TEMPO, $\text{TO}\cdot$) forming the products $\text{RCH}(\text{OT})\text{CO}_2\text{T}$ resulting from addition of one TEMPO radical to the carbonyl carbon and a second to the resulting radical. The rate constants estimated for TEMPO addition are the lowest so far determined for addition to ketenes, and together with previous results give a qualitative correlation of $\log k_2(\text{TEMPO}) = 1.20 \log k(\text{H}_2\text{O}) - 4.45$, $r = 0.94$ for the rate constants for hydration and for TEMPO reaction of the same ketenes. Computations using DFT methods for the reactions of $\text{CH}_2=\text{C}=\text{O}$ (**1**) and $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ with aminoxyl radicals are in agreement with the observed relative reactivities and selectivities.

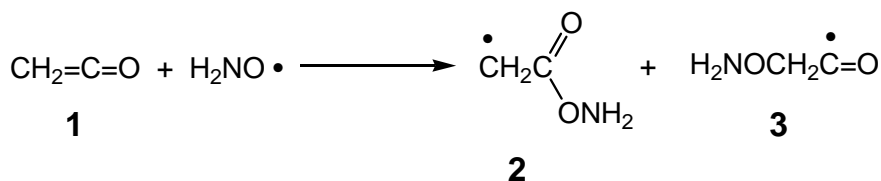
Keywords: Ketenes, TEMPO, aminoxyl radicals, reactivity correlation

Introduction

Previous studies in our laboratory have examined the reactivities of representative ketenes with H_2O ,¹ amines,² and the aminoxyl radical tetramethylpiperidinyloxy (TEMPO, $\text{TO}\cdot$).³ Very wide ranges of reactivity have been observed, with some parallels between the comparative rate constants of ketenes with H_2O as compared to amines. Rather surprisingly a semi-quantitative empirical correlation between the reactivities with water and TEMPO has been found as shown in eq. 1.^{3d,f}

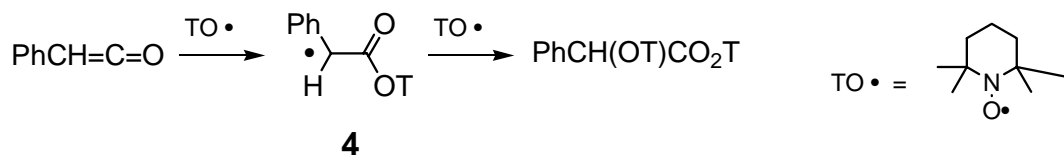
$$\log k_2(\text{TEMPO}) = 1.10 \log k(\text{H}_2\text{O}) - 3.79 \quad (1)$$

Free radical additions of $\text{H}\cdot$, $\text{CH}_3\cdot$, $\text{HO}\cdot$, $\text{F}\cdot$, $\text{Cl}\cdot$, and $\text{SiH}_3\cdot$ to ketene (**1**) are predicted by molecular orbital calculations to be highly exothermic.^{3a} Our earlier calculations for the stabilized aminoxyl radical $\text{H}_2\text{NO}\cdot$ using DFT methods indicated that attack at the carbonyl carbon forming **2** is exothermic, while attack at CH_2 forming **3** is endothermic, and less favored than attack at 1-C (carbonyl carbon atom) by approximately 25 kcal/mol (Scheme 1).^{3b}



Scheme 1

This prediction has been tested experimentally using tetramethylpiperidinyloxy (TEMPO, $\text{TO}\cdot$) with a variety of ketenes,^{3b-h} and as illustrated with phenylketene reaction occurs readily, with initial in-plane attack of TEMPO at the carbonyl carbon forming an α -acyl radical **4** which reacts further with a second TEMPO, forming a diaddition product (Scheme 2). With alkenyl, allenyl, and alkynylketenes allylic or propargylic rearrangement may occur.^{3c,e-g} Further computations of reactions of $\text{Cl}\cdot$ ^{4a} and $\text{OH}\cdot$ ^{4b} radicals with ketene have appeared, and there have been experimental studies of ketene reactions with $(\text{CF}_3)_2\text{NO}\cdot$,^{4c} $\text{NO}\cdot$,^{4d} $\text{NO}_2\cdot$,^{4d} and $\text{NO}_3\cdot$.^{4e}



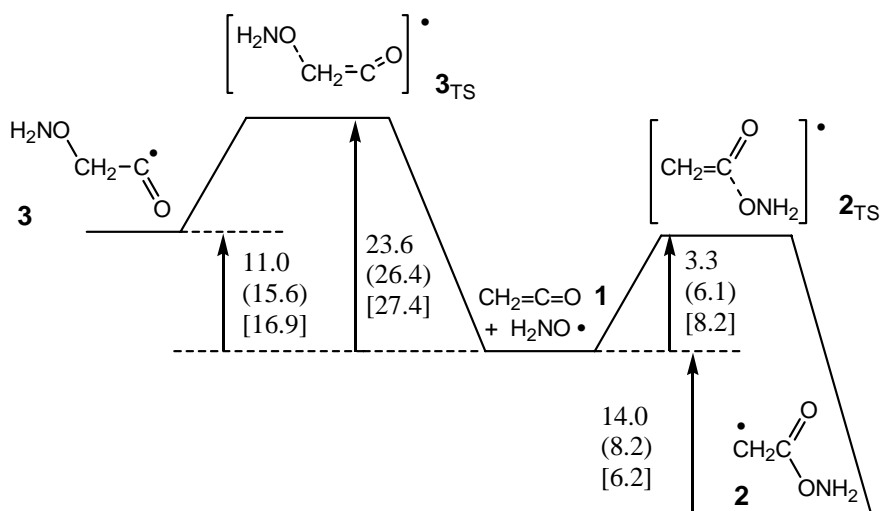
Scheme 2

The question remained whether ketene itself or [dimethyl(phenyl)silyl]ketene (**12**), which is even less reactive than ketene towards H_2O , would react with TEMPO. In 1936, Hurd and Williams^{4f} reported that reaction of $\text{Ph}_3\text{C}\cdot$ with $\text{CH}_2=\text{C}=\text{O}$ was not observed. We now report computations at higher levels, including transition structures, and experimental studies of radical reactions of the relatively unreactive substrates ketene (**1**) and dimethyl(phenyl)silyl]ketene (**12**).

Results

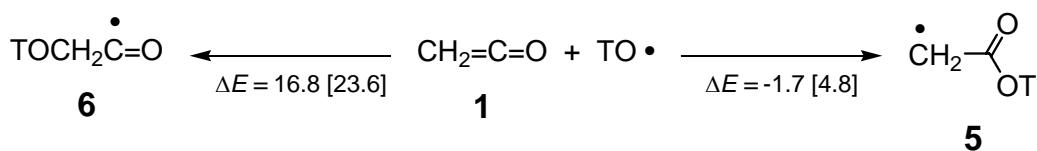
Computations of the reactions of $\text{CH}_2=\text{C}=\text{O}$ (**1**) with $\text{H}_2\text{NO}\cdot$ for addition at 1-C and at 2-C forming the respective transition structures $\mathbf{2}_{\text{TS}}$ and $\mathbf{3}_{\text{TS}}$ and products **2** and **3**, were carried out at the B3LYP/6-31G**/B3LYP/6-31G*, B3LYP/6-311G**//B3LYP/6-311G**, and B3LYP/6-

311++G**//B3LYP/6-311++G** levels, including zero point vibrational energies. The results are given in Table 1 and summarized in Scheme 3, and show that in agreement with our preliminary studies^{3b} that addition at 1-C is exothermic, and addition at 2-C is endothermic at all levels. Addition at 1-C is more favorable than addition at 2-C by 25.0, 23.8, and 23.1 kcal/mol, respectively, at increasingly higher levels, and additions at both positions are predicted to become less favorable at the higher levels. The transition structure **2**_{TS} for attack at 1-C is 20.3, 20.3, and 19.2 kcal/mol lower in energy than **3**_{TS} for attack at 2-C at increasingly higher levels, respectively. There are significant changes in the energies at different levels, but the energy differences are remarkably constant.



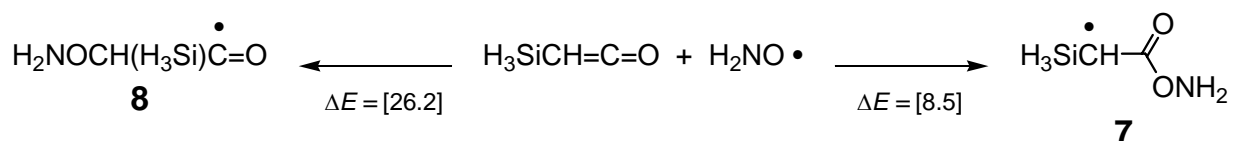
Scheme 3. Energy changes (kcal/mol) for H₂NO• addition to ketene at B3LYP/6-31G*, B3LYP/6-311G** (parentheses), and B3LYP/6-311++G** [brackets].

Computations for TEMPO addition to CH₂=C=O were carried out at the B3LYP/6-31G**//B3LYP/6-31G* and B3LYP/6-311++G**//B3LYP/6-311++G** levels, and the results are presented in Table 1 and summarized in Scheme 4. These show that addition to 2-C forming TOCH₂C(•)=O is less favorable than addition of H₂NO• at the same levels, by 5.8 and 6.6 kcal/mol, respectively, while addition at 1-C forming •CH₂CO₂T is less favorable by 11.2 and 11.0 kcal/mol, respectively (Scheme 4). Thus 1-C addition is favored by 18.5 and 18.8 kcal/mol at the two levels, respectively, or slightly less than found for addition of H₂NO•.



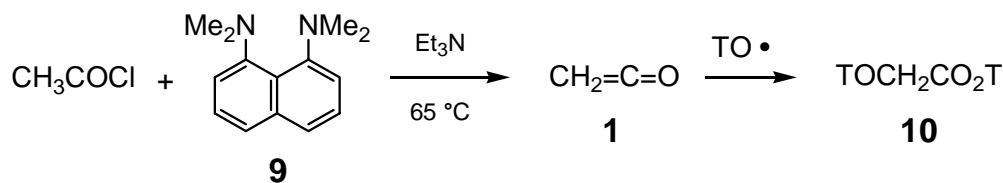
Scheme 4. Energy changes (kcal/mol) for TEMPO (TO•) addition at B3LYP/6-31G* and B3LYP/6-311++G** [brackets].

Reaction of $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ with $\text{H}_2\text{NO}\cdot$ was studied at the B3LYP/6-311++G**//B3LYP-311++G** level, and it was found that addition to 2-C forming $\text{H}_2\text{NOCH}(\text{H}_3\text{Si})\dot{\text{C}}=\text{O}$ is endothermic by 26.2 kcal/mol, while addition at 1-C forming $\cdot\text{CH}_2\text{CO}_2\text{NH}_2$ is endothermic by 8.5 kcal/mol (Scheme 5). Thus 1-C addition is favored by 17.7 kcal/mol, which is 5.5 kcal/mol less than the difference of 23.2 kcal/mol found for addition to $\text{CH}_2=\text{C}=\text{O}$ (**1**) at the same level (Scheme 1). Addition to $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ is less favorable compared to $\text{CH}_2=\text{C}=\text{O}$ (**1**) at 1-C by 9.2 kcal/mol, and by 14.7 kcal/mol at 1-C.



Scheme 5. Energy changes (kcal/mol) for $\text{H}_2\text{NO}\cdot$ addition to $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ at B3LYP/6-311++G** [brackets].

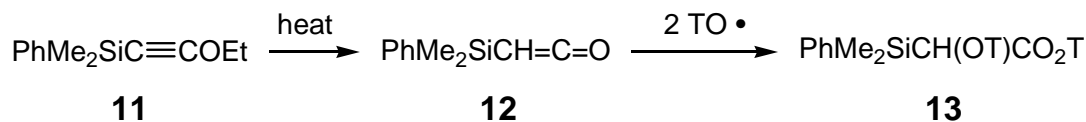
In experimental studies we find acetyl chloride generates ketene over time when reacted with 1,8-bis(dimethylamino)naphthalene (**9**) with a catalytic amount of triethylamine in toluene at 65 °C by the procedure of Lectka, et al.^{5a} as observed in situ by the IR absorption at 2135 cm^{-1} . When the reaction was carried out in the presence of TEMPO the adduct **10** was isolated in low yield (Scheme 6). For kinetic studies ketene (**1**) was prepared by pyrolysis (550 °C) of diketene,^{5b} and a stock solution in hexane was added to TEMPO in isooctane, and the decrease in the absorption of the ketene band was monitored at 2139 cm^{-1} . Rate constants for the reaction of ketene (**1**) at 25 °C (Table 2) were measured over a concentration range of TEMPO from 0.280 to 0.971 M, and gave a linear correlation $k(\text{obs}) = (7.62 \pm 0.66) \times 10^{-5} \text{ M}^{-1} \text{ s}^{-1} - (1.54 \pm 0.44) \text{ s}^{-1}$.



Scheme 6

Pyrolysis of alkyne **11** in a gas chromatograph gave dimethyl(phenyl)silyl]ketene (**12**) (Scheme 7).^{5c} Heating **12** in hexane with 6 equiv. of TEMPO gave a product mixture estimated by ^1H NMR to contain unreacted **12** and TEMPO adduct **13** in a 3:4 ratio, and possibly $\text{PhMe}_2\text{SiCHCO}_2\text{H}$ (**14**). Column chromatography gave **13** in a 47% yield based on consumption of **12**. The product **13** from addition of TEMPO was identified by its spectral properties, including the ^1H NMR which showed the 2 diastereotopic methyl groups attached to silicon as very sharp singlets at δ 0.50 and 0.52, and all eight methyl groups of the piperidinyl moieties as singlets. Careful examination of the product mixture showed only traces of $\text{TOCH}_2\text{CO}_2\text{T}$ (**10**).

Rate constants for reaction of dimethyl(phenyl)silyl]ketene (**12**) with TEMPO at 92.1 °C were measured over a range of TEMPO concentration of 0.40 to 1.20 M, and gave a linear correlation $k(\text{obs}) = (4.41 \pm 0.35) \times 10^{-5} \text{ M}^{-1} \text{ s}^{-1} - (0.39 \pm 0.29) \times 10^{-5} \text{ s}^{-1}$. To estimate a rate constant for **12** at 25 °C the same temperature dependence of the reactivity as found for the bisketene **14**^{3h} (Table 2) was assumed to apply, and this gave a rate constant for **12** of $4.9 \times 10^{-7} \text{ M}^{-1} \text{ s}^{-1}$.



Scheme 7

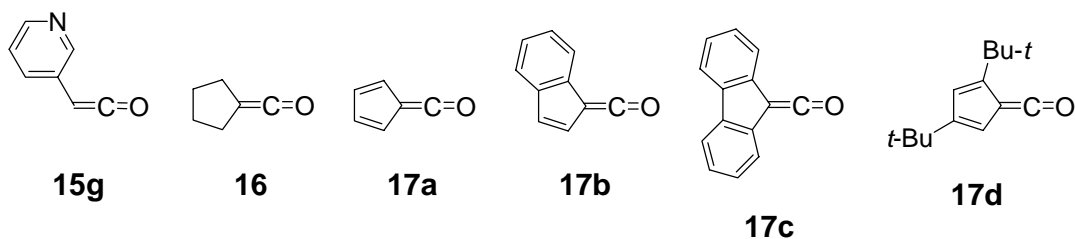
The experimental range that could be studied for **1** and **12** was limited by the low reactivity of both substrates and the volatility of ketene, and so data over a wider range of concentrations of TEMPO and at different temperatures was not obtained. A hint of possible curvature could be inferred in the plots of $k(\text{obs})$ vs [TEMPO], but until better experimental techniques are devised the present data will suffice for comparative purposes.

Discussion

Our previous theoretical^{3a,b} and experimental^{3b,h} studies have established the mechanism of TEMPO reaction with ketenes as involving initial in-plane radical attack on the carbonyl carbon forming acyl substituted radical intermediates, followed by attack of a second TEMPO forming the product (Scheme 2). The large accelerations on the rate of TEMPO attack by radical stabilizing aryl, vinyl, and alkynyl substituents (Table 2) confirm that this first step of attack on 1-C is rate limiting. The computational results and the occurrence of allylic and other rearrangements of the intermediate radicals in certain examples confirm attack on the carbonyl carbon is the first step. Simple acyl substituted radicals are known to have spin density predominantly on 2-C.^{3a}

The new computational studies confirm our previous conclusion^{3b} that attack of $\text{H}_2\text{NO}\cdot$ at 1-C of ketene is predicted to be favored, while attack at 2-C is strongly endothermic, and is more than 20 kcal/mol less favorable than attack at 1-C (Scheme 3). The barrier for attack at 1-C is also much lower than for attack at 2-C. Addition to either position is indicated to become progressively less favorable at higher levels, but the preference for attack at 1-C is large at all levels. This preference parallels that found previously for reaction of ketene with $\text{H}_2\text{N}\cdot$, $\text{HO}\cdot$, $\text{F}\cdot$, and $\text{Cl}\cdot$, a result that was attributed to the high conjugative stability of the amide, carboxylic acid, and acyl halide systems formed.^{3a} Addition of TEMPO to $\text{CH}_2=\text{C}=\text{O}$ (**1**) is predicted to be less favorable than for $\text{H}_2\text{NO}\cdot$, but attack at 1-C is still strongly preferred.

Rate constants for the reactions of ketenes with TEMPO and with water are collected in Table 2. Ketene (**1**) is found to be much less reactive than the ketenes **15a-d,f,g** (Table 2), all of which possess conjugating substituents which are expected to stabilize an intermediate from TEMPO addition to 1-C, which would have radical character on 2-C, as shown for **4** (Scheme 2). Ketene is also less reactive by a factor of 39 towards TEMPO than is cyclopentylidene-methanone **16**, in which the radical intermediate is stabilized by two alkyl groups. Also included in Table 2 are 3-pyridylketene (**15g**)^{2e} and the fulvenones **17**^{3f}.



Dimethyl(phenyl)silylketene (**12**) is even less reactive than ketene (**1**) towards TEMPO, by a factor of 1.6×10^3 . This is in agreement with the computational results, in which the reaction of $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ with $\text{H}_2\text{NO}\cdot$ at 1-C is predicted to be less favorable than for $\text{CH}_2=\text{C}=\text{O}$ by 14.7 kcal/mol (Table 1). The low reactivity of **12** may be attributed to the ground state stabilization of this ketene by the β -silicon effect, which is the favorable in-plane interaction of the C-Si bond with the electron deficient carbonyl group. This stabilization is lost in the intermediate **7** from addition of the aminoxyl radical at 1-C. The decrease in the stabilization of the intermediate radical due to the SiH_3 substituent is predicted to be 9.2 kcal/mol for addition at 2-C, compared to 14.7 kcal/mol for addition at 1-C, and this difference of 5.5 kcal/mol is indicative of the stabilization of the radical **8** formed by addition at 2-C by the β -silicon effect.

The computations at the B3LYP/6-311++G**//B3LYP/6-311++G** level predict that the additions at 1-C of TEMPO to $\text{CH}_2=\text{C}=\text{O}$ (**1**) (Scheme 4) and of $\text{H}_2\text{NO}\cdot$ to $\text{H}_3\text{SiCH}=\text{C}=\text{O}$ (Scheme 5) will be endothermic, by 4.8 and 8.5 kcal/mol, respectively. While these predictions may not be accurate for the solution phase reactions (Schemes 6, 7) such endothermic steps are not unreasonable, because addition of the second TEMPO as in Scheme 2 will occur at a near diffusion controlled rate, and so the initial attack would still be rate limiting.

The fit of the second order rate constants for the reactions of ketenes with TEMPO with the rate constants for hydration of the same ketenes (Table 2) gives the correlation of eq. 2 as illustrated in Figure 1. A qualitative correlation is observed, $r = 0.94$, and such a correlation between a free radical and a polar reaction is perhaps surprising. However in both cases in-plane attack by oxygen on the carbonyl carbon of the ketene is involved, and the oxygen of nitroxyl radicals is often considered as being nucleophilic, and possessing some negative charge.^{7a} The $\text{p}K_a$ of TEMPO has been estimated as -5.5 ± 1 ,^{7b} indicating the nitroxyl is about 3 $\text{p}K_a$ units less basic than CH_3OH .

$$\log k_2(\text{TEMPO}) = 1.18 \log k(\text{H}_2\text{O}) - 4.39 \quad (2)$$

The range in reactivity of the TEMPO reactions is 10^9 , as compared to 9×10^6 for the H_2O reactions. A variety of effects might be involved to different degrees in these reactions, including steric effects of the bulky tetramethylpiperidinyloxy moiety, solvation effects, and aromaticity in the fulvenones. However the overall correlation shows somewhat surprisingly that despite the differences there is a reasonable correlation of the reactivities. Presumably the stabilities of the reactants play an important role, and the energetic changes in going to the transition states are similar.

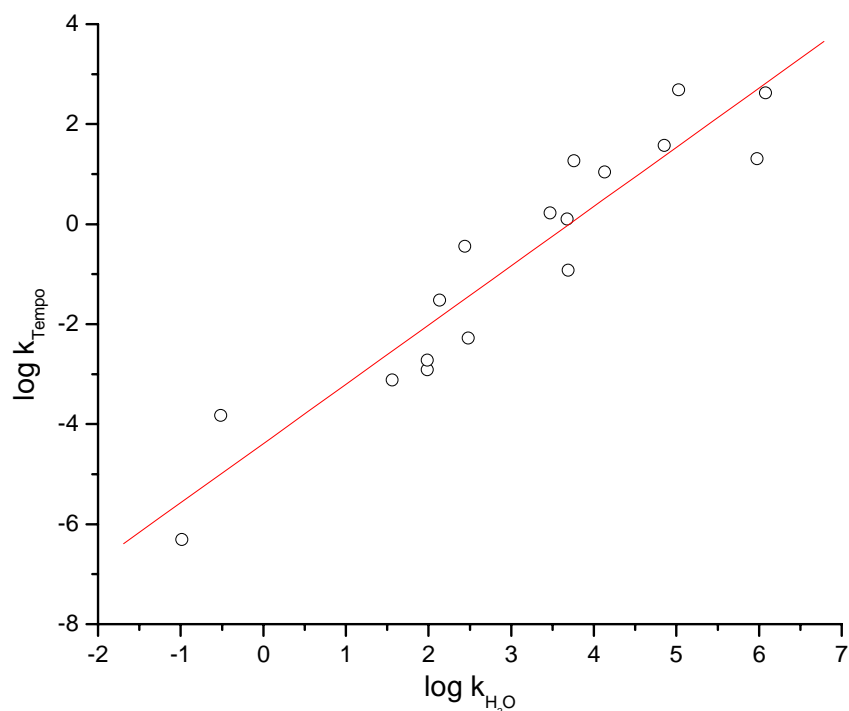


Figure 1. Correlation of rate constants k_2 for ketene reactions with TEMPO with rate constants for hydration.

Conclusions

In conclusion computational studies predict a strong preference for reactions of aminoxyl radicals with ketene at 1-C as compared to addition to 2-C, a preference that may be explained as due to the formation the stabilized ester function. Computations on $H_3SiCH=C=O$ indicate this ketene will be less reactive than $CH_2=C=O$ (**1**), as expected from the strong ground state stabilization of the reactant due to the β -silicon effect. Experimental studies of TEMPO reactions with $CH_2=C=O$ (**1**) and $Me_3SiCH=C=O$ agree with the computational predictions, and the rate

constants for the reactions of 17 ketenes with TEMPO give a surprisingly good correlation with the rate constants for reaction of the same ketenes with H₂O.

Table 1. Energies (Hartrees) for aminoxyl radical addition to ketene (**1**) and H₃SiCH=C=O^a

Structure	Level	<i>E</i>	ZPVE ^c	<i>E</i> + ZPVE	<i>E</i> (rel) [kcal/ mol]
H ₂ NO•	B3LYP/6-31G* ^b	-131.082450	0.025808	-131.056642	
CH ₂ =C=O		-152.598470	0.031747	-152.566723	
[H ₂ NO--CH ₂ =C=O]• (TS)		-283.647947	0.062191	-283.585756	23.6
H ₂ NOCH ₂ C(•)=O		-283.671200	0.065364	-283.605836	11.0
[CH ₂ =C(--ONH ₂)=O]• (TS)		-283.681099	0.062962	-283.618137	3.3
•CH ₂ CO ₂ NH ₂		-283.710760	0.065066	-283.645694	-14.0
H ₂ NO•	B3LYP/6-311G** ^b	-131.129766	0.026428	-131.103338	
CH ₂ =C=O		-152.647753	0.031577	-152.616176	
[H ₂ NO--CH ₂ =C=O]• (TS)		-283.739301	0.061888	-283.677413	26.4
H ₂ NOCH ₂ C(•)=O		-283.759077	0.064443	-283.694634	15.6
[CH ₂ =C(-ONH ₂)=O]• (TS)		-283.772393	0.062569	-283.709824	6.1
•CH ₂ CO ₂ NH ₂		-283.797270	0.064620	-283.732650	-8.2
H ₂ NO•	B3LYP/6-311++G**	-131.138719	0.025982	-131.112737	
CH ₂ =C=O		-152.651599	0.031510	-152.620089	
[H ₂ NO--CH ₂ =C=O]• (TS)		-283.750971	0.061883	-283.689088	27.4
H ₂ NOCH ₂ C(•)=O		-283.770160	0.064417	-283.705743	17.0
[CH ₂ =C(--ONH ₂)=O]• (TS)		-283.782118	0.062359	-283.719759	8.2
•CH ₂ CO ₂ NH ₂		-283.807236	0.064542	-283.742694	-6.2
TO•	B3LYP/6-31G*	-483.719770	0.263498	-483.456272	
CH ₂ =C=O		-152.598471	0.031745	-152.566726	
TOCH ₂ C(•)=O		-636.296871	0.300651	-635.996220	16.8
•CH ₂ CO ₂ T		-636.324928	0.299170	-636.025758	-1.7
TO•	B3LYP/6-311++G**	-483.855969	0.260681	-483.595288	
CH ₂ =C=O		-152.651598	0.031510	-152.620088	
TOCH ₂ C(•)=O		-636.474174	0.296329	-636.177845	23.6
•CH ₂ CO ₂ T		-636.502428	0.294705	-636.207723	4.8
H ₃ SiCH=C=O		-443.384353	0.048385	-443.335968	
H ₂ NOCH(SiH ₃)C(•)=O	-574.486280	0.079348	-574.406932	26.2	
•CH(SiH ₃)CO ₂ NH ₂	-574.514680	0.079479	-574.435201	8.5	

^a Using Gaussian 98 unless indicated. ^b Using Gaussian 94. ^c Zero point vibrational energy.

Table 2. Rate constants for the reaction of ketenes with TEMPO and with H₂O at 25 °C

	k_2 (TEMPO) (M ⁻¹ s ⁻¹)	k (H ₂ O) (s ⁻¹)	Ref.
CH ₂ =C=O (1)	7.61 x 10 ⁻⁴	36.5	this work
PhMe ₂ SiCH=C=O (12)	4.9 x 10 ⁻⁷	0.139	this work
(Me ₃ SiC=C=O) ₂ (14)	1.50 x 10 ⁻⁴	0.307	3h
Ph ₂ C=C=O (15a)	0.357	275	3h
PhCH=C=O (15b)	1.26	4.77 x 10 ³	3h
(<i>E</i>)-PhCH=CHCH=C=O (15c)	18.4	5.76 x 10 ³	3h
PhC≡CCH=C=O (15d)	37.2	7.16 x 10 ⁴	3h
<i>n</i> -BuCH=C=O (15e)	1.22 x 10 ⁻³	96.6	3h
1-(Naphthyl)CH=C=O (15f)	1.68	2.95 x 10 ³	3d
3-Pyridylketene (15g)	11.0	1.36 x 10 ⁴	3h
CyclopropylCH=C=O (15h)	5.26 x 10 ⁻³	304	3g
CH ₂ =CH(CH ₂) ₃ CH=C=O (15i)	1.89 x 10 ⁻³	96.6	3g
Cyclopentylideneketene (16)	2.98 x 10 ⁻²	230	3h
Pentafulvenone (17a)	20.2	9.00 x 10 ⁵	3f
Indenylideneketene (17b)	422	1.20 x 10 ⁶	3f
Fluorenylideneketene (17c)	482	1.07 x 10 ⁵	3f
2,4-Di-(<i>t</i> -Bu)pentafulvenone (17d)	0.120	4.91 x 10 ³	3f

Experimental Section

General Procedures. Reactions were conducted in dried glassware under an atmosphere of N₂ or Ar. Chromatography was carried out on silica gel. Triethylamine was dried by distillation from CaH₂, and Ar was bubbled through solutions for dehydrochlorinations for 1 h at the beginning of these reactions. Toluene was distilled from Na containing benzophenone. Solutions for TEMPO reactions were degassed with Ar.

Computations. Computations were carried out using Gaussian 94 with a Silicon Graphics Indigo computer and Gaussian 98 with a computer cluster from Velocet Communications, Toronto. Final energies include ZPVE, in contrast to those reported previously.^{3b}

TEMPO Addition to ketene (1) forming 2,2,6,6-tetramethyl-1-[2-oxo-2-[(2,2,6,6-tetramethylpiperidin-1-yl)oxy]ethoxy]piperidine (10). TEMPO (520 mg, 3.3 mmol), 1,8-bis(dimethylamino)naphthalene (370 mg, 1.7 mmol), and Et₃N (0.04 mL, 0.3 mmol) were added to a cooled solution of acetyl chloride (0.1 mL, 1.4 mmol) in 3 mL of toluene in a 25 mL round bottom flask. The solution was heated at 65 °C for 9 h, cooled and filtered, and the filtrate was extracted with 1N HCl and then with NaCl solution. The organic layer was dried and

concentrated, and chromatography (SiO₂, EtOAc/hexane 1:4) gave pure **10** (12 mg, 0.033 mmol, 2%), and in later fractions 8% of **10** containing some TEMPO. Pure oil **10** gave ¹H NMR (300 MHz, CDCl₃): δ 1.09 (s, 6H, 2 CH₃), 1.1–1.7 (m, 12H, 6 CH₂), 1.17 (s, 6H, 2 CH₃), 1.19 (s, 6H, 2 CH₃), 1.20 (s, 6H, 2 CH₃), 4.51 (s, 2H, CH₂OT). ¹³C NMR (100 MHz, CDCl₃): δ 17.17, 17.21, 20.2, 20.8, 29.9, 32.2, 33.0, 39.2, 39.8, 60.3, 60.4, 75.1 (C=O not observed). IR (CDCl₃): $\tilde{\nu}$ 1750, 1773 cm⁻¹. EIMS (70 eV): *m/z* 354 (0.1 M⁺), 198 (100, M⁺ –TO), 156 (31, TO⁺), 140 (99, T⁺). HR-EIMS: Calcd for C₂₀H₃₈N₂O₃ 354.2882, obs. 354.2872.

TEMPO Addition to phenyldimethylsilylketene (12) forming 1-[1-[dimethyl(phenyl)silyl]-2-oxo-2-[(2,2,6,6-tetramethylpiperidin-1-yl)oxy]ethoxy]-2,2,6,6-tetramethylpiperidine (13).

Ketene **12**^{5c} (27 mg, 0.15 mmol) and TEMPO (146 mg, 0.94 mmol) in 0.5 mL of hexanes in a sealed tube was heated at 75 °C for 20 h. The solvent was evaporated, and examination by ¹H NMR showed the presence of **13** and unreacted **12** in a 4:3 ratio, and possibly PhMe₂SiCH₂CO₂H. Chromatography (SiO₂, EtOAc/hexane 1:9) gave first 24 mg of almost pure material and then a fraction of pure oil **13** (21 mg, 0.043 mmol, 29%). ¹H NMR 300 MHz, CDCl₃) δ 0.50 (s, 3H, CH₃Si), 0.52 (s, 3H, CH₃Si), 0.54 (s, 3H, CH₃), 0.85 (s, 3H, CH₃), 1.04 (s, 3H, CH₃), 1.07 (s, 3H, CH₃), 1.14 (s, 3H, CH₃), 1.16 (s, 3H, CH₃), 1.25 (s, 3H, CH₃), 1.33 (s, 3H, CH₃), 1.0–1.7 (m, 12, 6H, CH₂), 4.74 (s, 1H, CHOT), 7.3–7.4 (m, 3H, Ph), 7.6–7.7 (m, 2H, Ph). ¹³C NMR 100 MHz, CDCl₃): δ –5.3, –3.6, 17.2, 17.5, 20.4, 20.8, 21.0, 31.9, 33.2, 33.9, 35.5, 39.6, 40.8, 41.4, 59.5, 60.0, 60.2, 61.9, 77.9, 128.0, 129.8, 134.6, 135.8, 171.2. IR (CDCl₃): $\tilde{\nu}$ 1760 cm⁻¹, 1772 cm⁻¹. EIMS (70 eV): *m/z* 488 (0.02 M⁺), 332 (19, M⁺ –TO), 176 (8, M⁺ –2 TO), 161 (63, M⁺ –2 TO, Me), 140 (100, T⁺). HR-EIMS: Calcd for C₂₈H₄₇N₂O₃Si 487.3356, obs. 487.3359.

Kinetics of TEMPO addition to ketene 1. Ketene (**1**) prepared by heating of diketene to 550 °C was condensed in hexane and stored at –78 °C. The solution showed the ketene IR band at 2139 cm⁻¹ and based on the UV absorbance of 0.04 at 325 nm (reported ε = 10)⁷ had a concentration of 0.5 M. An aliquot of this solution (60 μL) was injected into a solution of freshly sublimed TEMPO in isooctane (0.280 to 0.971 M) and the decrease in the ketene absorption with time was monitored at 2139 cm⁻¹.

Kinetics of TEMPO addition to phenyldimethylsilylketene (12). Aliquots (200 μL) of solutions containing ketene **12** (4 mL, 6.24 x 10⁻³ M) and freshly sublimed TEMPO (0.401 to 1.197 M) were sealed in glass ampoules and heated at 92.0 °C. Samples were removed at intervals and the decrease in the ketene absorption with time at 2112 cm⁻¹ was measured. These were fit to a first order rate expression and the derived rate constants gave a dependence on [TEMPO] $k(\text{obs}) = (4.41 \pm 0.35) \times 10^{-5} \text{ s}^{-1} \text{ M}^{-1} [\text{TEMPO}] - (3.89 \pm 2.94) \times 10^{-6} \text{ s}^{-1}$.

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