

## An unprecedented photodimerization of an anthraceno[2.2]paracyclophane

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**Dedicated to our friend and colleague, Jim Coxon, on the occasion of his 65<sup>th</sup> birthday**

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### Abstract

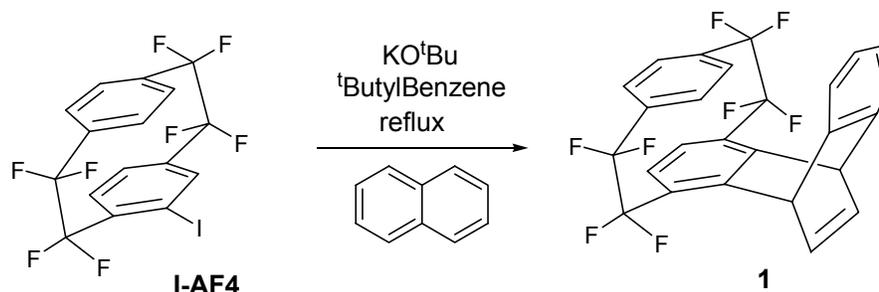
Octafluoro[2.2](1,4)anthraceno-paracyclophane slowly undergoes a clean photodimerization upon being exposed, in solution, to fluorescent light. X-ray crystal analysis of this photodimer indicated that the aromaticity of the anthracene ring directly involved in the paracyclophane system was observed to be destroyed by forming a cyclobutane ring, with the two molecules linked to each other in a head to tail fashion.

**Keywords:** Photochemistry, [2.2]paracyclophane, anthracene, X-ray crystal analysis

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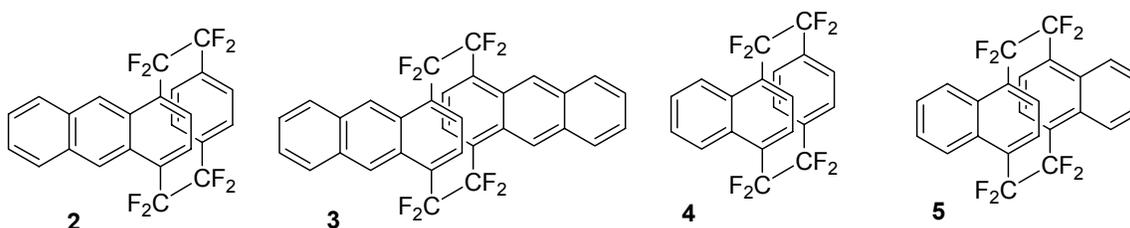
### Introduction

As a result of work in our laboratories during the last few years, it has been established that the aryne derived from treatment of 4-iodo-1,1,2,2,9,9,10,10-octafluoro[2.2]paracyclophane (**I-AF4**)<sup>1</sup> with potassium *t*-butoxide in a non-protic solvent has extraordinary Diels-Alder reactivity with various aromatic substrates, as exemplified by its efficient reaction with naphthalene, shown in Scheme 1.<sup>2</sup>

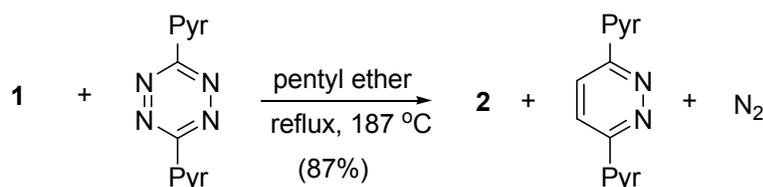


### Scheme 1

It was subsequently demonstrated that such Diels-Alder adducts could be used to prepare novel naphthaleno- and anthraceno-[2.2]phanes (**2-5** in Scheme 2) via their one-pot, three step reaction with 3,6-di-(2-pyridinyl)-1,2,4,5-tetrazine, as exemplified in Scheme 3 for the synthesis of [2.2](1,4)anthraceno-paracyclophane, **2**.<sup>3</sup> The process involves first a Diels-Alder cycloaddition to the tetrazine, then loss of N<sub>2</sub>, and lastly a retro-Diels-Alder reaction to give the product.



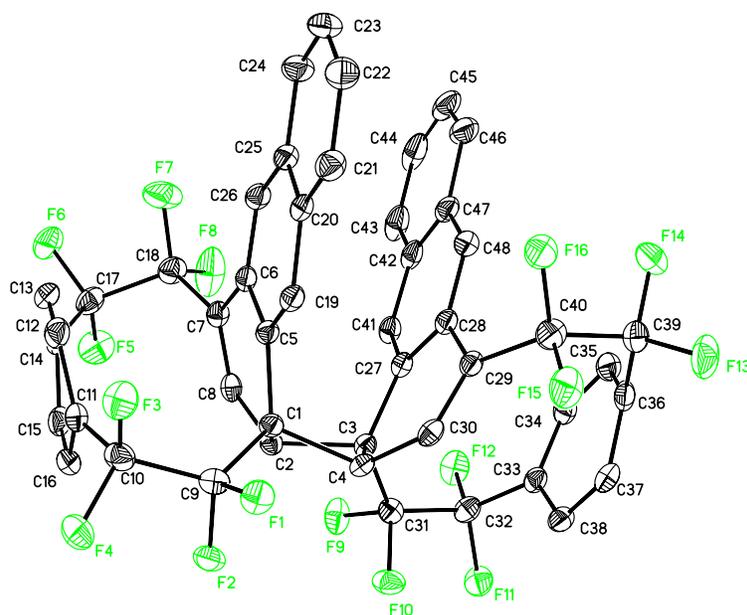
### Scheme 2



### Scheme 3

During the course of the preparation, purification and characterization of this series of naphthaleno- and anthracenophanes, it was found that **2** exhibited an unexpected lability that was uncharacteristic of the other members of the prepared series. When an NMR sample of pure **2** in CDCl<sub>3</sub> was allowed to stand for one week at room temperature, it was found that 30% of **2** had been converted to a new product, which could be separated from **2** by column chromatography.



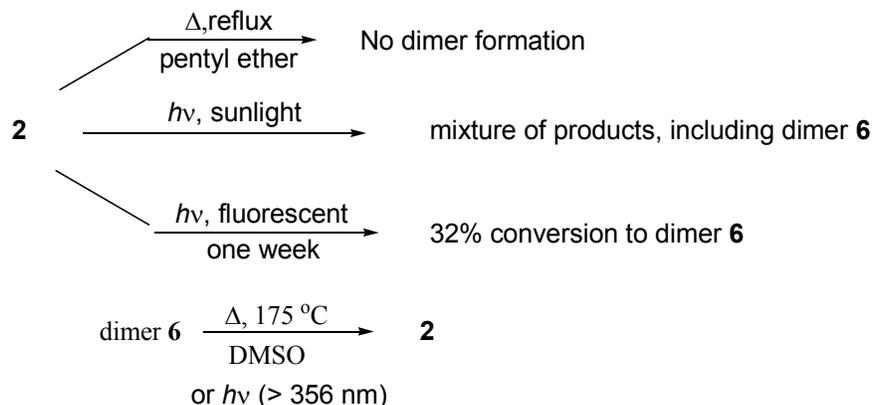


**Figure 1.** ORTEP drawing of dimer **6**.

The X-ray diffraction analysis of the compound derived from **2** indicated that there were two molecules of **6** in its asymmetric unit.<sup>4</sup> The aromaticity of the anthracene ring directly involved in the paracyclophane system was observed to be destroyed, with the two entities linked to each other in a head to tail fashion via a [2 + 2] cycloaddition. For molecule A, the formed cyclobutane ring had a twist angle of 31.8 ° and a torsion angle C1-C2-C3-C4 of 22.4 ° to relieve the steric strain of the two bulky anthracene moieties. The two fluorine bridges C1-C9-C10-C11 and C14-C17-C18-C7 had torsion angles 40.1 ° and 26.9 °, respectively, in molecule A. The anthracene moiety in bonds C5-C1-C2-C8 and C27-C3-C4-C30 had torsion angles 29.5 ° and 33.2 °, respectively. The cyclobutane ring of molecule B had a twist angle of 23.5 ° with torsion angles of 35.1 ° and 20.6 ° in the two fluorine bridges.

**Dimer formation and chemistry.** In principle, the dimer could have been formed by either a thermal or a photochemical [2 + 2] process. However, heating **2** at reflux in pentyl ether overnight in the dark did not lead to any significant dimer formation. Dimer **6** was itself thermally stable at temperatures up to 150 °C. Above this temperature, **6** was observed to cleave cleanly back to monomer **2** with a rate constant of  $9.8 \times 10^{-5} \text{sec}^{-1}$  at 175 °C, which corresponds to a half life of 118 minutes. Thus dimerization cannot be a thermal process.

On the other hand, subjecting a chloroform solution of **2** to bright sunlight for one day led to some loss of **2** with formation of a complex mixture of products, which included dimer **6**. When such a chloroform solution of **2** was exposed to *simple fluorescent lighting* for one week at room temperature, a clean 32% conversion to dimer **6** was observed (Scheme 5). If the dimer is subjected to *uv light* (>356 nm), a reversion to monomer **2** is observed, thus indicating a great sensitivity of the photoequilibrium of this system to the wavelength of light.



**Scheme 5.** Dimer formation and chemistry.

The photochemistry of anthracene and substituted derivatives of anthracene goes back to the work of Fritzsche in 1867, who observed the formation of an insoluble substance when solutions of anthracene were exposed to sunlight.<sup>5</sup> Like **6**, this substance was found to revert to anthracene upon heating (melting). Upon further examination, many years later, the product of this photoreaction was determined to be the symmetrical, “sandwich” dimer formed by linking both anthracenes at their 9, 10-positions.<sup>6,7</sup>

Anthraceno-paracyclophane **2** is a pale yellow crystalline solid, and its unique photoreactivity likely derives from the fact that, with significant absorption extending well past 425 nm, its uv spectrum extends into the visible region, whereas the naphthalenophanes, **4** and **5**, are white solids with no absorption in the visible region, their absorptions extending only just past 350 nm. Anthracenophane **3** is also yellow and might well be expected to exhibit similar photochemistry. However, its limited solubility at room temperature in all solvents thus far tested has precluded examination of its bimolecular photochemistry. If an appropriate solvent can be found that will allow such a study, it will be reported in the future.

The dimer itself is a *white* solid, thus explaining the photoequilibrium that is favorable to its formation from **2** in visible light.

Since the uv spectra of the fluorinated and non-fluorinated anthracenophanes do not differ significantly, it is expected that the hydrocarbon system analogous to **2**<sup>8-10</sup> should exhibit analogous photodimerization reactivity.

Future work will also involve examination of the photochemistry of the two naphthalenophanes, **4** and **5**, to see if analogous photodimerization processes may be observed for these two systems when using uv light.

## Experimental Section

**Dimer 6.** mp 150 °(dec)(turns yellow and slowly reverts to monomer **2**);  $^1\text{H}$  NMR,  $\delta$  7.60(d,  $J=8.1\text{Hz}$ , 2H), 7.55(d,  $J=8.9\text{Hz}$ , 2H), 7.40(ddd,  $J=8.0$ , 6.9, 1.2Hz, 4H), 7.33(ddd,  $J=8.0$ , 6.8, 1.1Hz, 2H), 7.18(d,  $J=7.7\text{Hz}$ , 2H), 6.74(dd,  $J=3.1$ , 1.5Hz, 2H), 6.36(d,  $J=8.5\text{Hz}$ , 2H), 6.31(s, 2H), 6.02(d,  $J=8.4\text{Hz}$ , 2H), 5.70(d,  $J=8.3\text{Hz}$ , 2H), 4.21(d,  $J=7.0\text{Hz}$ , 2H);  $^{13}\text{C}$  NMR,  $\delta$  134.3, 133.7, 132.4, 131.9, 131.2, 129.8, 129.5, 127.9, 127.6, 129.8, 128.1, 122.7, 128.8, 128.4, 131.6, 128.4, 126.1, 118.1, 118.0, 116.8, 115.0, 50.8, 45.0, 43.2; MS (EI):  $m/z$  452(21), 276(100), 176(49); HRMS (ESI-FT-ICR) Calc for  $\text{C}_{48}\text{H}_{24}\text{F}_{16} + \text{Na}$  927.1515, found 927.1504; UV ( $\text{CH}_2\text{Cl}_2$ ):  $\lambda_{\text{max}}$  ( $\log \epsilon$ ) = 358.1 (3.22), 339.8 (3.64), 323.8 (3.97), 270.6 (4.55) nm.

**X-ray experimental.** Data for dimer **6** were collected at 173 K on a Siemens SMART PLATFORM equipped with a CCD area detector and a graphite monochromator utilizing  $\text{MoK}_\alpha$  radiation ( $\lambda = 0.71073 \text{ \AA}$ ). Cell parameters were refined using up to 8192 reflections. A full sphere of data (1850 frames) was collected using the  $\omega$ -scan method ( $0.3^\circ$  frame width). The first 50 frames were re-measured at the end of data collection to monitor instrument and crystal stability (maximum correction on I was  $< 1 \%$ ). Absorption corrections by integration were applied based on measured indexed crystal faces.

The structure was solved by the Direct Methods in *SHELXTL6*,<sup>11</sup> and refined using full-matrix least squares. The non-H atoms were treated anisotropically, whereas the hydrogen atoms were calculated in ideal positions and were riding on their respective carbon atoms. The asymmetric unit consists of two chemically equivalent but crystallographically independent molecules, one molecule of dichloromethane in general position and another half disordered in two positions, in the asymmetric unit, around an inversion center. The latter half molecule could not be modeled properly, thus program SQUEEZE,<sup>12</sup> a part of the PLATON<sup>13</sup> package of crystallographic software, was used to calculate the solvent disorder area and remove its contribution to the overall intensity data. A total of 1180 parameters were refined in the final cycle of refinement using 25923 reflections with  $I > 2\sigma(I)$  to yield  $R_1$  and  $wR_2$  of 4.15% and 9.00%, respectively. Refinement was done using  $F^2$ .<sup>4</sup>

CIF file containing X-ray structural data for dimer **6**.

## Acknowledgements

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