

Synthesis and some stereoselective radical additions of bis[(phenyldimethylsilyl)methyl]tin dihydride

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Dedicated to Professor Rosa Lederkremer on her 70th birthday

Abstract

The synthesis of the new bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) starting from chloromethyldimethylsilicon chloride is reported. A study on the radical additions of **6** to various mono- and disubstituted alkynes shows that these reactions take place with high and in some cases with complete stereoselectivity. Full ¹H-, ¹³C-, and ¹¹⁹Sn-NMR data of the new organotins are given.

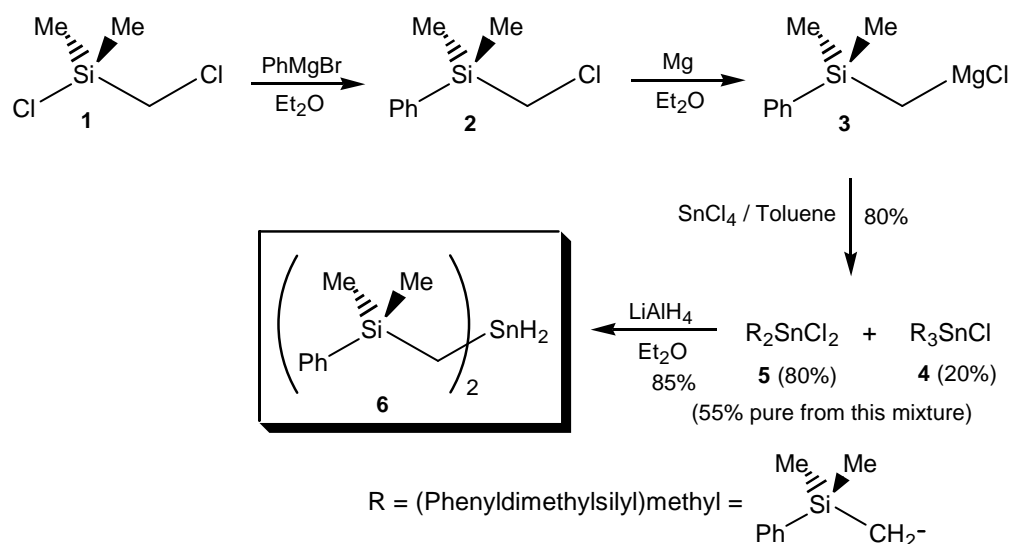
Keywords: Bis[(phenyldimethylsilyl)methyl]tin dihydride, stereoselective radical additions, divinyltins

Introduction

Organotin hydrides have found many applications in organic synthesis not only as reducing reagents but also as intermediates in the generation of carbon-carbon bonds, and for the preparation of compounds such as vinylstannanes, which are invaluable starting materials for cross-coupling reactions.¹ In previous studies we have shown that the size of the organic ligands attached to the tin atom affects not only the reactivity but also the stereoselectivity of the reactions of these compounds.² We have also recently reported the synthesis of tin compounds containing the ligand (phenyldimethylsilyl)methyl like tris[(phenyldimethylsilyl)methyl]tin dihydride.³ Taking into account these facts, and following our investigations on the relationship between the steric volume of the substituents attached to the tin atom and the reactivity as well as the stereoselectivity of organotin hydrides, we considered it to be of interest to carry out the synthesis of bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) and to study some reactions of this new bulky tin dihydride.

Results and Discussion

The total synthesis of bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) was carried out according to Scheme 1.

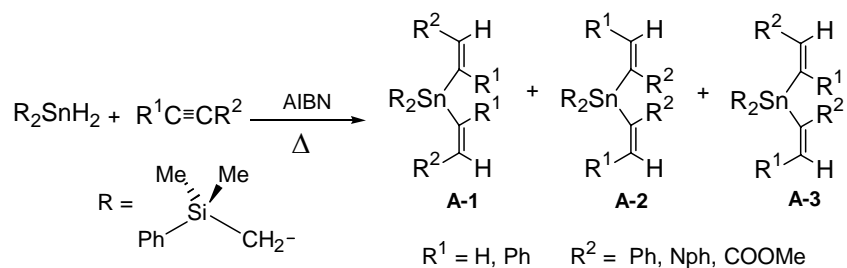


Scheme 1

The alkylation of tin tetrachloride with (phenyldimethylsilyl)methylmagnesium chloride (**3**) using a ratio Grignard reagent/SnCl₄ = 2.1 was repeated eight times, always obtaining a mixture of chlorides **4** and **5** in almost equal proportions, with an average yield of about 80%. These mixtures were very soluble in methylene chloride, diethyl ether, hexane, and petroleum ether, even at low temperatures (0 to -23 °C), and we were unable to recrystallize **5** pure from these mixtures. Taking into account these results, we decided to carry out a study of the reaction under different experimental conditions and **3**/SnCl₄ ratios. In all the experiments the Grignard reagent was added dropwise to the solution of tin tetrachloride in toluene. The results obtained are summarized in Table 1. From the mixture obtained in entry 7 (Table 1), an average of ten experiments, we were able to obtain dichloride **5** pure in an average yield of 55% by crystallization from heptane. It should be noted that we have previously obtained **5** but in very low yield.³ The reduction of **5** with lithium aluminum hydride gave the desired bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) in 85% yield.

To test whether it was possible to perform stereoselective hydrostannations under radical conditions using dihydride **6**, we carried out the addition of this hydride to phenylethyne (**a**), 2-naphthylethyne (**b**), methyl propiolate (**c**), diphenylethyne (**d**), and methyl 3-phenylpropiolate (**e**). Taking into account previous works,^{2f, 4} we considered it possible that under these conditions the main products of the hydrostannations should be those corresponding to an *anti* addition of the hydride to the carbon-carbon triple bond. Thus, the *anti* addition of tin dihydride **6** to these alkynes could lead to the three possible regioisomers **A-1/A-3** shown in Scheme 2.

A summary of the obtained results as well as the best reaction conditions is shown in Table 2.



Scheme 2

Table 1. Tin tetrachloride alkylation under different conditions

Entry	3/SnCl ₄ ratio	Reaction Conditions ^a	Yield (%)	Product Relationships (%) ^b	
				R ₂ SnCl ₂	R ₃ SnCl
1	2.1	(1) rt; (2) reffx 3 h; (3) rt 15 h	78	51	49
2	1.7	(1) rt; (2) reffx 3 h; (3) rt 15 h	80	54	46
3	2.1	(1) rt; (2) rt 15 h	76	62	38
4	2.0	(1) rt; (2) rt 48 h	82	55	45
5	2.0	(1) 0°C; (2) rt 2 h	79	72	28
6	1.6	(1) 0°C; (2) reffx 1h	78	70	30
7	1.5	(1) 0°C; (2) rt 2 h	80	80	20

^a All the reactions were carried out by adding the Grignard reagent (GR) in diethyl ether onto a solution of SnCl₄ in toluene; rt = room temperature; reffx = reflux. ^b From ¹¹⁹Sn NMR spectra.

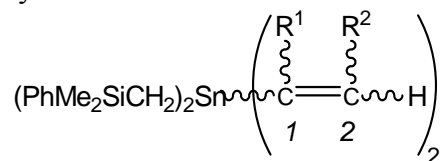
Table 2. Radical additions of bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) to alkynes^a

Compd N ^o	Alkyne	Time (h)	R ¹	R ²	A-1 (%)	A-2 (%)	A-3 (%)	Yield ^b (%)	¹¹⁹ Sn (ppm) ^c
7	a	4	H	Ph	100	---	---	86	-91.6
8	b	4	H	2-Nph	100	---	---	85	-91.0
9 & 10	c	1	H	COOMe	50 (9)	---	30 (10)	62 ^d	9 : -102.7 10 : -86.9
11	d	4	Ph	Ph	100	---	---	90	-80.3
12 & 13	e	1	Ph	COOMe	80 (12)	20 (13)	---	78	12 : -79.4 13 : -94.3

^a The reactions were carried out under a nitrogen atmosphere; ratio alkyne/hydride **6** = 2.5; AIBN 0.01 equiv; without solvent; at 75 °C. ^b Yields of products isolated from chromatography. ^c In CDCl₃; in ppm with respect to Me₄Sn. ^d The formation of some product of monoaddition was also detected.

In Tables 3 and 4 are collected ^{13}C - and ^1H -NMR data of the new organotin **7-13**.

Table 3. ^{13}C NMR data of divinyltin adducts **7-13**^a

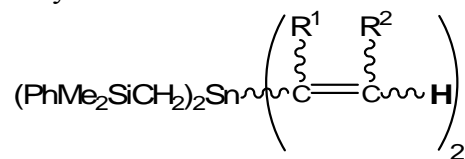


Adduct N ^o	R ¹	R ²	C(1) [¹ J(Sn,C)]	C(2) [² J(Sn,C)]	R ^{1b} [² J(Sn,C)]	R ^{2b} [³ J(Sn,C)]	Other signals
7	H	Ph	134.31 (456.3)	146.70 (nd)	---	140.61 (33.0)	c
8	H	2-Nph	134.99 (452.6)	146.64 (nd)	---	138.08 (31.4)	d
9	H	COOMe	160.47 (416.1)	133.31 (11.4)	---	168.13 (29.6)	e
10	H	COOMe	157.03 (433.4)	134.56 (10.5)	---	168.36 (30.9)	f
	COOMe	H	148.65 (390.6)	141.51 (23.3)	170.63 (43.9)	---	f
11	Ph	Ph	149.97 (420.4)	144.16 (24.0)	139.54 (24.3)	146.78 (35.1)	g
12	Ph	COOMe	140.53 (447.8)	154.05 (12.1)	137.92 (22.3)	171.11 (42.0)	h
13	COOMe	Ph	140.04 (422.1)	153.93 (30.2)	170.84 (nd)	137.17 (36.3)	i

^a In CDCl_3 ; chemical shifts, δ , in ppm with respect to TMS; $^n\text{J}(\text{Sn,C})$ and $^1\text{J}(\text{Si,C})$ coupling constants, in Hz (in parentheses); nd = not determined; Nph = Naphthyl. ^b Chemical shifts of the carbons attached to C(1) and C(2). ^c -2.67 (275.0) (49.8); 0.00 (53.6) (12.5); 127.50; 127.68; 128.21; 128.69; 133.41; 133.44; 141.36 (20.8). ^d -2.46 (277.5) (69.8); -0.00 (53.6) (13.0); 125.92; 126.20; 126.45; 127.66; 127.70; 127.73; 128.12; 128.65; 132.93; 133.33; 133.47; 133.51; 141.24 (22.8). ^e -2.57 (341.5) (51.5); 0.00 (52.0) (15.7); 51.51; 127.48; 128.47; 133.33; 141.57 (23.1). ^f -2.46 (333.7) (48.8); 0.00 (51.7) (14.5); 0.18 (52.4) (16.5); 51.78; 51.96; 127.75; 128.79; 133.53; 139.60 (15.6). ^g -0.73 (268.3) (47.3); 0.00 (52.1) (12.3); 125.93; 127.24; 127.51; 127.80; 127.88; 128.08; 128.36; 128.52; 133.26; 141.31 (24.3). ^h 0.00 (52.4) (13.4); 1.38 (307.8) (49.2); 52.06; 127.96; 128.53; 129.02; 129.25; 133.72; 141.66 (24.6). ⁱ -1.81 (324.3) (54.0); 0.00 (51.4) (14.2); 51.56; 127.14; 127.75; 128.07; 128.49; 133.00; 141.30 (20.1).

The stereochemistry assigned to the new compounds **7-13** follows from the large $^3J(\text{Sn,H})$ coupling constants, between 109 and 165 Hz (Table 4), that indicate the existence of *trans* H-C=C-Sn linkages in these compounds.

Table 4. ^1H NMR data of divinyltin adducts **7-13**



Adduct N°	Chemical shifts (δ , in ppm) ^a
7	0.00 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 71.2]; 0.25 (CH ₃) (s, 12H); 6.12 (R ¹ = H) [d, 2H, $^3J(\text{H,H})$ 13.5; $^2J(\text{Sn,H})$ 72.0]; 7.27-7.48 (Ph & R ² = Ph) (m, 20H); 7.53 (H) [d, 2H, $^3J(\text{H,H})$ 13.5; $^3J(\text{Sn,H})$ 165.6].
8	-0.20 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 75.8]; -0.00 (CH ₃) (s, 12H); 6.02 (R ¹ = H) [d, 2H, $^3J(\text{H,H})$ 12.7; $^2J(\text{Sn,H})$ 72.5]; 7.03-7.71 (Ph & R ² = 2-Nph) (m, 24H); 7.53 (H) [d, 2H, $^3J(\text{H,H})$ 12.7; $^3J(\text{Sn,H})$ 163.9].
9	0.00 (CH ₃) (s, 12H); 0.18 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 86.3]; 3.52 (R ² = COOMe) (s, 6H); 6.39 (H) [d, 2H, $^3J(\text{H,H})$ 12.6; $^3J(\text{Sn,H})$ 136.5]; 6.81 (R ¹ = H) [d, 2H, $^3J(\text{H,H})$ 12.6; $^2J(\text{Sn,H})$ 67.9]; 7.07-7.27 (Ph) (m, 10H).
10	0.00 (CH ₃) (s, 6H); 0.01 (CH ₃) (s, 6H); 0.14 (SiCH ₂ Sn) [s, 2H, $^2J(\text{Sn,H})$ 84.7]; 0.16 (SiCH ₂ Sn) [s, 2H, $^2J(\text{Sn,H})$ 85.9]; 3.48 (R ¹ = COOMe) (s, 3H); 3.50 (R ² = COOMe) (s, 3H); 5.65 (R ² = H) [d, 1H, $^2J(\text{H,H})$ 2.5; $^3J(\text{Sn,H})$ 61.2]; 6.42 (H) [d, 1H, $^3J(\text{H,H})$ 12.1; $^3J(\text{Sn,H})$ 145.6]; 6.61 (H) [d, 1H, $^2J(\text{H,H})$ 2.5; $^3J(\text{Sn,H})$ 120.4]; 6.91 (R ¹ = H) [d, 1H, $^3J(\text{H,H})$ 12.1; $^2J(\text{Sn,H})$ 73.8]; 7.08-7.26 (Ph)(m, 10H).
11	-0.28 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 76.6]; 0.00 (CH ₃) (s, 12H); 6.85-7.22 (H , Ph, R ¹ = R ² = Ph) (m, 32H).
12	-0.06 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 82.2]; 0.00 (CH ₃) (s, 12H); 3.48 (R ² = COOMe) (s, 6H); (Ph & R ¹ = Ph) 6.94-7.20 (m, 20H); 8.00 (H) [s, 2H, $^3J(\text{Sn,H})$ 133.9].
13	-0.07 (SiCH ₂ Sn) [s, 4H, $^2J(\text{Sn,H})$ 75.0]; 0.00 (CH ₃) (s, 12H); 3.43 (R ¹ = COOMe) (s, 6H); 7.05-7.32 (Ph & R ² = Ph) (m, 20H); 8.21 (H) [s, 2H, $^3J(\text{Sn,H})$ 109.4].

^a In CDCl₃; chemical shifts, δ , in ppm with respect to TMS; $^nJ(\text{Sn,H})$ and $^nJ(\text{H,H})$ coupling constants, in Hz (in parentheses).

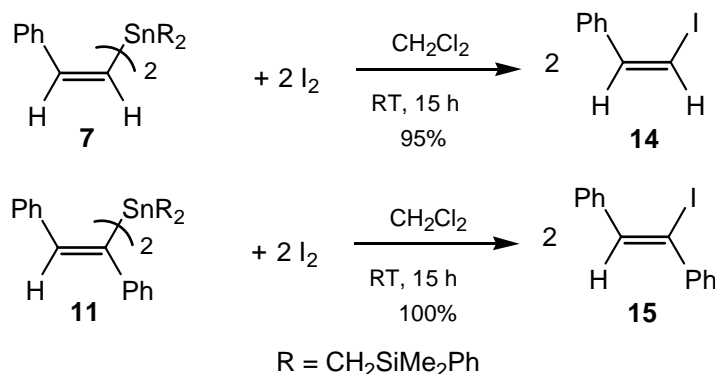
Table 2 shows that the addition of hydride **6** to phenylethyne, naphthylethyne, methyl propiolate, diphenylethyne, and methyl 3-phenylpropiolate leads exclusively to the *Z*-vinylstannanes resulting from an *anti* attack, typical of radical hydrostannations.^{2f,3,4}

A remarkable fact is the formation of the *Z*-adducts as sole products in the additions of hydride **6** to phenylethyne, naphthylethyne, and diphenylethyne. In previous studies, we have

shown that the addition of trineophyl- and tris[(phenyldimethylsilyl)methyl]tin hydrides to phenylethyne and diphenylethyne^{2f,3} leads to the *E*-adducts as the sole products. We explained the formation of the *E*-adducts by considering that the isomerization of the initially formed kinetic *Z*-products by further addition/elimination of the stannyl radical would favor, under the reaction conditions, the formation of the thermodynamically more stable *E*-vinylstannanes. The exclusive formation of the *Z*-adducts in the case of the addition of hydride **6** could be explained by taking into account that this organotin dihydride might be more reactive than the bulky trineophyl- and tris[(phenyldimethylsilyl)methyl]tin hydrides. Therefore, it should be capable of effecting a much faster hydrogen transfer to the intermediate alkyl radical during the second step of the radical propagation chain leading straight to the products.

Although in most cases symmetrical, adducts were obtained, i.e., both vinyl ligands with the same regiochemistry, an interesting fact is that in one of the five hydrostannations carried out with dihydride **6** we obtained as a minor product a mixed adduct, i.e., an adduct containing as vinyl ligands the two possible regioisomers: divinylstannane **10**.

The chemical reactivity of the new divinylstannanes is similar to that of other monovinyltriorganotins. Thus, the iododestannylation of adducts **7** and **11** leads to the corresponding iodovinyl compounds **14** and **15** in very high yields and with complete retention of the *Z* configuration, as shown in Scheme 3.



Scheme 3

Conclusions

Our results demonstrate that by an adequate selection of the experimental parameters it is possible to obtain pure bis[(phenyldimethylsilyl)methyl]tin dichloride (**5**) in 55% yield from the alkylation of tin tetrachloride with (phenyldimethylsilyl)methylmagnesium chloride (**3**). The reduction of **5** with lithium aluminum hydride leads to bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) with 85% yield. The hydrostannylation under radical conditions of internal and terminal alkynes with **6** takes place stereoselectively leading to the products corresponding to an

anti addition of the hydride to the carbon-carbon triple bond. The chemical reactivity of the new divinylstannanes is similar to that of other monovinyltriorganotinins.

Experimental Section

General Procedures. NMR spectra were obtained in a Bruker ARX 300 instrument. Infrared spectra were recorded with a Nicolet Nexus FT spectrometer. Mass spectra were obtained using a Finnigan MAT Model 8230 at Dortmund University (Germany). Microanalyses were performed at Dortmund University. The melting points were determined in a Kofler hot stage and are uncorrected. Dimethyl(chloromethyl)chlorosilane (**1**) was purchased and (chloromethyl)phenyldimethylsilane (**2**) was obtained following known procedures.⁵ All the solvents and reagents used were analytical reagent grade.

Synthesis of bis[(phenyldimethylsilyl)methyl]tin dichloride (**5**)

To magnesium turnings (1.12 g, 0.046 mol) in dry diethylether (10 ml) under stirring, was added slowly and dropwise a solution of (chloromethyl)phenyldimethylsilane (**2**) (7.78 g, 0.042 mol) in diethyl ether (30 ml). The preparation was carried out under an atmosphere of nitrogen. The reaction was exothermic and after the addition was completed the mixture was refluxed during 1 h. Then the reaction mixture was stirred for 15 h at room temperature. The solution of the Grignard reagent **3** thus obtained had a 1.04 M concentration.

To a solution of SnCl₄ (3.3 ml, 7.29 g, 0.028 mol) in dry toluene (66 ml) at 0°C under nitrogen, was added slowly and dropwise with stirring a solution of the Grignard reagent **3** in diethyl ether (0.042 mol, 40 ml of 1.04 M solution). The mixture was stirred for 2 h at room temperature. After cooling at 0° C, the reaction mixture was decomposed with a saturated solution of NH₄Cl (*ca* 50 ml). The organic layer was separated and the aqueous layer was extracted three times with diethyl ether (*ca* 30 ml each). The combined organic extracts were dried over anhydrous MgSO₄ and the solvent was removed under reduced pressure. The ¹¹⁹Sn NMR of the resulting solid (11.45 g) showed it to consists of a mixture of bis[(phenyldimethylsilyl)methyl]tin dichloride (**5**) and tris[(phenyldimethylsilyl)methyl]tin chloride (**4**) and in a ratio $5/4 = 4:1$ (11.45 g, 80 % yield of organotin compounds based upon SnCl₄ used). Repeated recrystallizations of the resulting mixture from n-heptane afforded **5** as a white crystalline solid (4.81 g, 0.010 mol, 55%). Mp: 75-76 °C. It should be noted that we have previously reported the synthesis of **5** but in only a 21% yield.³ In reference 3 could be found the spectroscopic characteristics of **5**.

Synthesis of bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**)

To a suspension of LiAlH₄ (0.36 g, 9.4 mmol) in dry diethyl ether (20 ml) under atmosphere of argon at room temperature, was added dropwise a solution of bis[(phenyldimethylsilyl)methyl]tin dichloride (**5**) (3.10 g, 6.3 mmol) in dry diethyl ether (30 mL). The mixture was heated under reflux for 5 h and a saturated solution of ammonium chloride was then added (*ca* 15

ml). The organic layer was separated, and the aqueous was extracted three times with diethyl ether. The combined organic extracts were dried over anhydrous MgSO_4 . Removal of the solvent under reduced pressure gave **6** as a liquid (2.2 g, 5.4 mmol, 85%). IR (KBr, cm^{-1}): $\nu_{\text{Sn-H}}$ 1833. $^1\text{H-NMR}$ (C_6D_6 , Me_4Si) δ ppm: 0.00 [t, 4H, $^3\text{J}(\text{H}, \text{H}) = 2.5$ Hz, $^2\text{J}(\text{Sn}, \text{H}) = 76.6$ Hz]; 0.34 (s, 12 H); 4.79 [q, 2 H, $^3\text{J}(\text{H}, \text{H}) = 2.5$ Hz, $^1\text{J}(\text{Sn}, \text{H}) = 1804.7$ Hz]; 7.24-7.55 (m, 10H). $^{13}\text{C-NMR}$ (C_6D_6 , Me_4Si) δ ppm: -8.19 [$^1\text{J}(\text{Sn}, \text{C}) = 270.8$ Hz, $^1\text{J}(\text{Si}, \text{C}) = 48.6$ Hz]; 0.00 [$^3\text{J}(\text{Sn}, \text{C}) = 58.1$ Hz, $^1\text{J}(\text{Si}, \text{C}) = 16.5$ Hz]; 128.70; 129.79; 134.29; 141.09 [$^3\text{J}(\text{Sn}, \text{C}) = 18.1$ Hz]. $^{119}\text{Sn-NMR}$ (C_6D_6 , Me_4Sn) δ ppm: -207.9. Anal. Calcd. for $\text{C}_{18}\text{H}_{28}\text{Si}_2\text{Sn}$: C, 51.56; H, 6.73. Found: C, 51.80; H, 6.80.

Addition of bis[(phenyldimethylsilyl)methyl]tin dihydride (**6**) to substituted alkynes under radical conditions. Typical Procedure

A mixture of phenylacetylene (0.12 g, 1.19 mmol), dihydride **6** (0.20 g, 0.48 mmol), and AIBN as a catalyst, was heated under argon at 75 °C during 4 h (this optimal time of reaction and temperature was indicated by earlier experiments in which the reaction was monitored by taking samples at intervals and observing the disappearance of the Sn-H absorption by IR). Column chromatography (silica gel 60) of the crude product afforded compound **7** (0.26 g, 0.41 mmol, 86 %) in the fraction eluted with hexane- diethyl ether (99:1).

Bis[(phenyldimethylsilyl)methyl]di[(Z)-styryl]stannane (7). MS (m/z , rel. int.): 520 (8%, $[\text{M}-\text{C}_8\text{H}_7]^+$, Sn-pattern); 474 (100%, $[\text{M}-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 417 (7%, $[\text{M}-2(\text{C}_8\text{H}_7)]^+$, Sn-pattern); 371 (23%, $[\text{M}-(\text{C}_8\text{H}_7)-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 269 (21%, $[\text{SnCH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$, Sn-pattern); 223 (12%, $[\text{M}-(\text{C}_8\text{H}_7)-2(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 135 (27%, $[\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 120 (13%, $[\text{Sn}]^+$, Sn-pattern); 103 (6%, $[\text{C}_8\text{H}_7]^+$). Anal. Calcd. for $\text{C}_{34}\text{H}_{40}\text{Si}_2\text{Sn}$: C, 65.49; H, 6.46. Found: C, 65.70; H, 6.60.

Bis[(phenyldimethylsilyl)methyl]di[(Z)-(2-naphthalen-2-yl)vinyl]stannane (8). MS (m/z , rel. int.): 574 (100%, $[\text{M}-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 420 (88%, $[\text{M}-(\text{C}_{12}\text{H}_9)-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 306 (11%, $[\text{M}-2(\text{C}_{12}\text{H}_9)]^+$, Sn-pattern); 273 (20%, $[\text{SnC}_{12}\text{H}_9]^+$, Sn-pattern); 269 (49%, $[\text{SnCH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$, Sn-pattern); 153 (34%, $[\text{C}_{12}\text{H}_9]^+$); 135 (40%, $[\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 120 (17%, $[\text{Sn}]^+$, Sn-pattern). Anal. Calcd. for $\text{C}_{42}\text{H}_{44}\text{Si}_2\text{Sn}$: C, 69.70; H, 6.13. Found: C, 69.80; H, 6.20.

Bis[(phenyldimethylsilyl)methyl]di[(Z)-methyl 3-acrylate]stannane (9). MS (m/z , rel. int.): 503 (21%, $[\text{M}-\text{C}_4\text{H}_5\text{O}_2]^+$, Sn-pattern); 439 (100%, $[\text{M}-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 269 (7%, $[\text{SnCH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$, Sn-pattern); 205 (15%, $[\text{SnC}_4\text{H}_5\text{O}_2]^+$, Sn-pattern); 149 (5%, $[\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 135 (20%, $[\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 120 (9%, $[\text{Sn}]^+$, Sn-pattern); 85 (8%, $[\text{C}_4\text{H}_5\text{O}_2]^+$). Anal. Calcd. for $\text{C}_{26}\text{H}_{36}\text{O}_4\text{Si}_2\text{Sn}$: C, 53.16; H, 6.17. Found: C, 59.70; H, 7.00.

Bis[(phenyldimethylsilyl)methyl][(Z)-methyl 3-acrylate](methyl 2-acrylate)stannane (10). MS (m/z , rel. int.): 503 (22%, $[\text{M}-\text{C}_4\text{H}_5\text{O}_2]^+$, Sn-pattern); 439 (100%, $[\text{M}-(\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph})]^+$, Sn-pattern); 269 (10%, $[\text{SnCH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$, Sn-pattern); 205 (10%, $[\text{SnC}_4\text{H}_5\text{O}_2]^+$, Sn-pattern); 149 (5%, $[\text{CH}_2\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 135 (22%, $[\text{Si}(\text{CH}_3)_2\text{Ph}]^+$); 120 (11%, $[\text{Sn}]^+$, Sn-pattern); 85 (3%, $[\text{C}_4\text{H}_5\text{O}_2]^+$). Anal. Calcd. for $\text{C}_{26}\text{H}_{36}\text{O}_4\text{Si}_2\text{Sn}$: C, 53.16; H, 6.17. Found: C, 59.80; H, 7.10.

Bis[(phenyldimethylsilyl)methyl]di[(Z)-1,2-diphenylvinyl]stannane (11). MS (*m/z*, rel. int.): 627 (15%, [M-(CH₂Si(CH₃)₂Ph)]⁺, Sn-pattern); 597 (100%, [M-C₁₄H₁₁]⁺, Sn-pattern); 419 (10%, [Sn(CH₂Si(CH₃)₂Ph)₂]⁺, Sn-pattern); 269 (15%, [SnCH₂Si(CH₃)₂Ph]⁺, Sn-pattern); 179 (51%, [C₁₄H₁₁]⁺); 135 (27%, [Si(CH₃)₂Ph]⁺); 120 (6%, [Sn]⁺, Sn-pattern). Anal. Calcd. for C₄₆H₄₈Si₂Sn: C, 71.22; H, 6.23. Found: C, 70.90; H, 6.10.

Bis[(phenyldimethylsilyl)methyl]di[(Z)-(3-phenyl)methyl 3-acrylate]stannane (12). MS (*m/z*, rel. int.): 591 (100%, [M-(CH₂Si(CH₃)₂Ph)]⁺, Sn-pattern); 579 (14%, [M-C₁₀H₉O₂]⁺, Sn-pattern); 430 (10%, [M-(CH₂Si(CH₃)₂Ph)-(C₁₀H₉O₂)]⁺, Sn-pattern); 281 (5%, [SnC₁₀H₉O₂]⁺, Sn-pattern); 269 (8%, [SnCH₂Si(CH₃)₂Ph]⁺, Sn-pattern); 149 (10%, [CH₂Si(CH₃)₂Ph]⁺); 135 (14%, [Si(CH₃)₂Ph]⁺); 120 (13%, [Sn]⁺, Sn-pattern); 91 (8%, [C₇H₇]⁺). Anal. Calcd. for C₃₈H₄₄O₄Si₂Sn: C, 61.71; H, 5.99. Found: C, 61.70; H, 5.80.

Bis[(phenyldimethylsilyl)methyl]di[(Z)-(3-phenyl)methyl 2-acrylate]stannane (13). MS (*m/z*, rel. int.): 591 (100%, [M-(CH₂Si(CH₃)₂Ph)]⁺, Sn-pattern); 579 (10%, [M-C₁₀H₉O₂]⁺, Sn-pattern); 430 (12%, [M-(CH₂Si(CH₃)₂Ph)-(C₁₀H₉O₂)]⁺, Sn-pattern); 281 (8%, [SnC₁₀H₉O₂]⁺, Sn-pattern); 269 (10%, [SnCH₂Si(CH₃)₂Ph]⁺, Sn-pattern); 149 (10%, [CH₂Si(CH₃)₂Ph]⁺); 135 (12%, [Si(CH₃)₂Ph]⁺); 120 (9%, [Sn]⁺, Sn-pattern); 91 (11%, [C₇H₇]⁺). Anal. Calcd. for C₃₈H₄₄O₄Si₂Sn: C, 61.71; H, 5.99. Found: C, 61.50; H, 5.90.

Iododestannylation reactions: synthesis of (Z)-1-iodo-2-phenylethene (14) and (Z)-1-iodo-1,2-diphenylethene (15)

To a solution of compound **7** (0.25 g, 0.40 mmol) in dry CH₂Cl₂ (4.5 ml) under argon was added iodine (0.22 g, 0.88 mmol). The mixture was stirred at room temperature for 15 h with monitoring of the reaction by TLC. Column chromatography on silica gel 60 (relation 70:1) of the crude product gave the iodovinyl compound **14** in the fraction eluted with hexane (0.17 g, 0.76 mmol, 95%). Similarly, iododestannylation of bis[(phenyldimethylsilyl)methyl]di[(Z)-1,2-diphenylvinyl]stannane (**11**) leads quantitatively to compound **15**.

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